

GURU NANAK COLLEGE (AUTONOMOUS)

(Affiliated to University of Madras and Re-Accredited at 'A' Grade by NAAC)

Guru Nanak Salai, Velachery, Chennai – 600042.



B.Sc. Chemistry

(SEMESTER PATTERN WITH CHOICE BASED CREDIT SYSTEM)

Syllabus

(For the candidates admitted in the Academic year 2019-20 and thereafter)

Vision

- ▣ To enhance the quality of education beyond the text book / syllabi based – exam oriented system to research and analytical based learning.
- ▣ To encourage the learners of exceptional quality to take up research and motivate them to contribute to the needs of the society.
- ▣ To encourage the faculty to constantly involve themselves in research in addition to the regular work, which would enable them to develop research oriented learning skills.

Mission

- ▣ To inculcate the scientific methodology of learning chemistry by focusing more on practicals.
- ▣ To enhance the creativity in learning chemistry among the learners using visual aids.
- ▣ To circulate the scientific methodology of learning among the learners.
- ▣ To facilitate the learning of chemistry in a systematic manner that would enrich the creativity among the learners.
- ▣ To produce quality graduates and post graduates to excel in the field of education / research / industry.
- ▣ To produce and to modernise the infra structure to impart and understand the importance of practical skill accuracy and data interpretation.
- ▣ To encourage the learners to participate in the teaching – learning process to enhance their analytical and problem solving skill and to develop leadership qualities.
- ▣ To motivate the students by conducting seminars/workshops with the inputs of eminent scientists, distinguished alumni and industrialist.
- ▣ Visit to Industries and scientific center to have exposure on sophisticated instruments and recent developments in chemistry.

PROGRAMME OUTCOME

After completion of the programme, the student will be able to

- PO 1: demonstrate the concepts and theories relating to the branches of chemical sciences such as organic, inorganic, physical and analytical chemistry.
- PO 2: utilize skills in problem solving, critical thinking, analytical reasoning in chemistry domain and use modern experimental techniques
- PO 3: develop a creative scientific mind to communicate effectively the scientific ideas and their impact on socio-economic issues and sensitize the need for a green environment
- PO 4: handle chemical materials safely by taking into account their physical and chemical properties including any specific hazards associated with their use.
- PO 5: apply the knowledge of chemistry to function effectively as an entrepreneur in chemical or related industries.

PROGRAMME SPECIFIC OUTCOME

The students at the time of graduation will be able to

- PSO 1: establish their skill in planning and conducting chemistry experiments, enabling them to handle classes at the secondary level.
- PSO 2: develop a creative scientific presentation to the society based on the need and communicate effectively the scientific ideas and their impact on socio-economic issues.

B.Sc., - CHEMISTRY
COURSE STRUCTURE (2019-22) Batch

Semester	Part	Course Component	Subject Code	Name of the course	Credits	Hours	Internal	External	Total
Semester I	I	Language	19UTAM141/ 19UHIN141	Language I	3	6	50	50	100
	II	English	19UENG241	English I	3	4	50	50	100
	III	Core Paper-I	19UCHE301	Basic Chemistry – I	4	5	50	50	100
		Core Practical - I	19UCHE302P	Volumetric Analysis and Inorganic Preparations	*	3	*	*	*
		Allied - I	19UMAT331	Mathematics – I	5	8	50	50	100
	IV	NME**	19UNME401E	Dairy Chemistry	2	2	50	50	100
		Soft Skill	19UGSL401	Introduction to study skills	3	2	50	50	100
Total Credits: 20 / Total Hours per week: 30									
Semester II	I	Language	19UTAM142/ 19UHIN142	Language - II	3	6	50	50	100
	II	English	19UENG242	English - II	3	4	50	50	100
	III	Core Paper-II	19UCHE303	Basic Chemistry - II	4	5	50	50	100
		Core Practical - II	19UCHE302P	Volumetric Analysis and Inorganic Preparation	3	3	50	50	100
		Allied - II	19UMAT332	Mathematics - II	5	8	50	50	100
	IV	NME**	19UNME402E	Food Chemistry	2	2	50	50	100
		Soft Skill	19UGSL402	Life skills	3	2	50	50	100
Total Credits: 23 / Total Hours per week: 30									
Semester III	I	Language	19UTAM143/ 19UHIN143	Language – III	3	6	50	50	100
	II	English	19UENG243	English – III	3	4	50	50	100
	III	Core Paper-III	19UCHE304	Organo-Oxygen Compounds	4	6	50	50	100
		Core Practical - III	19UCHE305P	Inorganic Qualitative Analysis	*	3	*	*	100
		Allied –III	19UPHY331	Physics - I	3	6	50	50	100
	IV	Allied Practical	19UPHY332P	Allied Physics Practical - I	*	3	*	*	100
	Soft Skill	19UGSL403	Job oriented Skills	3	2	50	50	100	
Total Credits: 16 / Total Hours per week: 30									

Semester	Part	Course Component	Subject Code	Name of the course	Credits	Hours	Internal	External	Total
Semester IV	I	Language	19UTAM144/ 19UHIN144	Language – IV	3	6	50	50	100
	II	English	19UENG244	English – IV	3	4	50	50	100
	III	Core Paper-IV	19UCHE306	Chemistry of s and p Block Elements	4	5	50	50	100
		Core Practical - II	19UCHE305P	Inorganic Qualitative Analysis	3	3	50	50	100
		Allied - II	19UPHY333	Physics – II	3	5	50	50	100
	IV	Allied Practical	19UPHY332P	Allied Physics Practical – I	4	3	50	50	100
			19UEVS401	EVS	2	2	50	50	100
		Soft Skill	19UGSL404	Computing Skills	3	2	50	50	100
Total Credits: 25 / Total Hours per week: 30									
Semester V	III	Core Paper-V	19UCHE307	Organo Nitrogen Compounds and Natural Products	4	4	50	50	100
		Core Paper-VI	19UCHE308	Chemical Kinetics and Electrochemistry	4	4	50	50	100
		Core Paper-VII	19UCHE309	Special topics in Chemistry	4	4	50	50	100
		Core Paper-VIII	19UCHE310	Analytical Chemistry - I	4	4	50	50	100
		Core Practical - III	19UCHE311P	Gravimetric Analysis	*	3	*	*	*
		Core Practical - IV	19UCHE312P	Analysis and Preparation of Organic Compounds	*	2	*	*	*
		Core Practical - V	19UCHE313P	Physical Chemistry Practical	*	3	*	*	*
	Elective – I [#]	19UCHE317	IDE [#] – Chemistry in Everyday life	5	5	50	50	100	
	IV	Internship [#]		Internship [#]	2				
			Value Education	2	1	50	50	100	
Total Credits: 25 / Total Hours per week: 30									
Semester VI	III	Core Paper-IX	19UCHE314	Thermodynamics and solutions	4	4	50	50	100
		Core Paper-X	19UCHE315	Chemistry of d & f block elements and Coordination chemistry	4	4	50	50	100
		Core Paper-XI	19UCHE316	Analytical Chemistry - II	4	4	50	50	100
		Elective – II ^{***}	19UCHE318	Industrial Chemistry	5	4	50	50	100
		Elective – III ^{***}	19UCHE319	Polymer Chemistry	5	4	50	50	100
		Core Practical -III	19UCHE311P	Gravimetric Analysis	4	4	50	50	100
		Core Practical - IV	19UCHE312P	Analysis and preparations of Organic Compounds	3	3	50	50	100
	Core Practical - V	19UCHE313P	Physical Chemistry Practical	3	3	50	50	100	
IV			Extension activities	1					
Total Credits: 33 / Total Hours per week: 30									
Grand Total Credits: 142/ Total Hours per week: 180									

- * The Practical Examinations will be conducted at the end of even semester.
- ** The Students must choose one NME paper in semester – I and one NME paper in semester – II.
- *** The students must choose Two Elective papers in Semester - VI from the list of offered electives.
- # The students must choose Inter Disciplinary Elective (IDE) offered by other departments.
- ## The students must undergo summer internship for **two weeks** after the fourth semester and the report to be submitted.

SEMESTER - I

**CORE THEORY – I
BASIC CHEMISTRY-I**

SUBJECT CODE: 19UCHE301	THEORY	MARKS 100
SEMESTER: I	CREDITS: 4	TOTAL HOURS: 60

Objectives:

Introduction to fundamental concepts of Physical, Organic and Inorganic Chemistry.

UNIT - 1: Atomic Structure

(10 hrs)

Rutherford's atomic model. Bohr's theory of Hydrogen atom - Postulates, Bohr's radius, energy of hydrogen like species - Planck's quantum theory - Photoelectric effect - Compton effect- Hydrogen atomic spectrum - Particle and wave nature of electron - deBroglie equation - Heisenberg's uncertainty principle.

Quantum mechanical postulates, Operators - addition, subtraction, multiplication, linear, Hermitian, Commutator, Vector, Laplacian and Hamiltonian. Schrodinger's wave equation (Derivation not required). Significance of Ψ and Ψ^2 . Wave mechanical concept of atomic orbitals - Shapes of orbitals - Quantum numbers. Zeeman effect.

Pauli's exclusion principle, Hund's rule - its basis and applications.

UNIT - 2: Periodic Classification and Periodicity in Properties

(10 hrs)

Aufbau principle, effective nuclear charge, screening effect, Slater's rule - applications and limitations. Electronic configuration of elements- extra stability of half - filled and completely filled atomic orbitals.

Classification of elements - s, p, d and f block elements. Modern periodic table - position of hydrogen in the periodic table. Periodicity of the following properties - atomic radius, ionic radius, ionization potential, electron affinity and electro negativity, horizontal, vertical and diagonal relationships in the periodic table.

Scales of electronegativity - Pauling, Mulliken and Allred - Rochow.

UNIT - 3: Chemical Bonding

(15 hrs)

Types of bond - ionic bond – factors influencing the formation of ion in compounds -ionization energy, electron affinity and lattice energy. Determination of lattice energy – Born Haber cycle, Inert pair effect, Fajan's rule.

Covalent bond – polarity of covalent bond, percentage ionic character of covalent bond, dipole moment- Molecular structures of CO₂, H₂O, NH₃ and CH₄- Bond characteristics - bond length, bond strength, bond angle and bond energy.

Valence Bond theory – sigma (σ) and pi (π) bonds- Hybridisation, Valence shell electron pair repulsion theory (VSEPR) and geometries of molecules – BeCl₂, H₂O, BF₃, NH₃, XeF₄, BrF₃, PCl₅, SF₆ and IF₇.

Molecular Orbital Theory - Bonding and antibonding orbitals, bond order, applications of MO theory to H₂, He₂, N₂, O₂, O²⁺, O²⁻, HF and CO. Comparison between VBT and MO theory - Hydrogen bonding – types and consequences.

UNIT - 4: Nomenclature of Organic Compounds and Isomerism

(10 hrs)

Nomenclature of organic compounds: IUPAC system of nomenclature of organic compounds – mono and bifunctional compounds.

Hybridisation and shapes of molecules - methane, ethane, ethylene, acetylene and benzene. Structural isomerism: chain isomerism, position isomerism, functional isomerism, metamerism and tautomerism.

UNIT - 5: Electronic Effects, Types of Reactions, Types of Intermediates (15 hrs)

Organic reactions – types (Substitution, Elimination, Addition, Rearrangement) with examples. Electron displacement effects - inductive, electromeric, mesomeric, resonance, hyperconjugation, steric effect and their applications (acid, base strength and dipole moment) Cleavage of bonds – homolytic and heterolytic fissions. Reactive intermediates - carbocations, carbanions, free radicals, carbenes, nitrenes and benzyne - their formation and stability.

TEXT BOOKS:

1. Arun Bahl and B. S. Bahl, Advanced Organic Chemistry, S. Chand & Company Pvt.Ltd, second edition, 2006.
2. M. K. Jain and S. C. Sharma, Modern Organic Chemistry, Vishal Publishing Co., fourth edition, 2003.
3. S. M. Mukherji and S. P. Singh, Reaction Mechanism in Organic Chemistry, Macmillan India Ltd, third edition, 1994.
4. B. R. Puri, L. R. Sharma and K. C. Kalia, Principles of Inorganic Chemistry, Milestone Publishers and Distributors, New Delhi, thirtieth edition, 2009.
5. P. L. Soni and Mohan Katyal, Textbook of Inorganic Chemistry, Sultan Chand & Sons, twentieth edition, 2006.

REFERENCE BOOKS:

1. J.E.Huheey, E.A. Keiter and R.L. Keiter, Inorganic Chemistry, Harper Collins, New York, fourth edition, 1993.
2. D.F. Shriver and P. W. Atkins, Inorganic Chemistry, W. H. Freeman and Co, London, fifth edition, 2010.
3. R. T. Morrison and R. N. Boyd, Organic Chemistry, Printice - Hall of India Limited, New Delhi, sixth edition, 1992.

Question paper pattern:

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

SEMESTER - II

**CORE THEORY – 2
BASIC CHEMISTRY-II**

SUBJECT CODE: 19UCHE303	THEORY	MARKS 100
SEMESTER: II	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To provide knowledge on fundamentals of chemistry.

UNIT –1: Nuclear chemistry **(15 hrs)**

Fundamental particles of the nucleus - nucleon - terminology, nucleides, isotopes, isobars, isotones, mirror nuclei, nuclear radius, nuclear mass and nuclear forces operating between the nucleons. n/p ratio, curves, stability belts. Nuclear binding energy, Mass defect, simple calculations involving mass defect and bond energy per nucleon. Shell model - Magic numbers - liquid drop model.

Radioactivity - group displacement law, natural radioactivity – radioactive series including neptunium series. Artificial radioactivity – induced radioactivity – uses of radio isotopes. Nuclear fission – Nuclear energy - Nuclear reactors – Breeder reactor – Nuclear fusion.

UNIT –2: Solid State **(10 hrs)**

Classification of solids, isotropic and anisotropic crystals, elements of symmetry, crystal systems, space lattices - bravais lattice, designation of planes, Miller indices, unit cell. Packing of ions in crystals (CCP, BCC and HCP).

X - ray diffraction – derivation of Bragg's equation, Discussion of structures of NaCl, CsCl and ZnS. Determination of Avogadro's number, Problems related to solid state chemistry. Defects in crystals - Frenkel and Schotky defects.

UNIT –3: Hydrocarbons

(10 hrs)

Alkanes – mechanism of free radical substitution in alkanes.
Cycloalkanes - general methods of preparation (Wurtz's reaction, Dieckmann's reaction) – Bayer's strain theory and theory of strainless rings.

Alkenes - preparation and properties – Electrophilic and free radical addition -Orientation of addition reaction (Markovnikov's rule and peroxide effect) – addition reaction, oxidation, ozonolysis and hydroboration. Hydroxylation with KMnO_4 , OsO_4 , allylic substitution by NBS. Diene – Classification – stability and reactivity of 1, 2 and 1, 4 - addition. Synthesis of diene - 1, 3 - butadiene, isoprene and chloroprene.

Alkynes – preparation and properties – acidity of alkynes, formation of acetylides, addition reaction, oxidation, ozonolysis and hydroboration.

UNIT –4: Principles of Quantitative and Qualitative Analysis

(10 hrs)

Volumetric analysis – Principles involved in acid - base, precipitation, complexometric and redox titrations – indicators and their choice. Definition and calculation of normality, molality, molarity, mole fraction and ppm.

Definition and examples of primary and secondary standards. Calculation of equivalent weights of acid, base, salt, metal, oxidizing and reducing agents.

Qualitative analysis – theory behind separation of groups - solubility product - common ion effect - Interfering anions and their removal. Brown ring test, Nessler's reagent, Prussian blue-Detection of phosphate, borate, fluoride, oxalate and chloride.

Unit-5: Aromatic Hydrocarbons, Nucleophilic Substitution and Elimination Reactions (15 hrs)

Aromaticity – Huckel’s rule with respect to benzene, naphthalene, anthracene, phenanthraene- Electrophilic substitution in aromatic compounds-general mechanism; nitration, sulphonation, halogenations, Friedel-Crafts alkylation and acylation. Orientation (directive influence) and reactivity in mono substituted benzenes.

Polynuclear hydrocarbons-naphthalene, anthracene, phenanthraene - preparation, properties, uses and carcinogenicity.

Nucleophilic substitution: S_N1 , S_N2 , S_{Ni} reactions-mechanisms. Elimination reactions: $E1$ and $E2$ reactions and mechanisms. Elimination Vs Substitution.

TEXT BOOKS

1. M. K. Jain, S. C. Sharma, Modern Organic Chemistry, Vishal Publishing, fourth reprint, 2003.
2. S.M. Mukherji and S.P. Singh, Reaction Mechanism in Organic Chemistry, Macmillan India Ltd., third edition, 1994.
3. B.R. Puri, L.R. Sharma and K.C. Kalia, Principles of Inorganic Chemistry, Milestone Publishers and Distributors, New Delhi, thirtieth edition, 2009.
4. J.D. Lee, Concise Inorganic Chemistry, Blackwell Science, fifth edition, 2005.
5. P.L. Soni and Mohan Katyal, Textbook of Inorganic Chemistry, SultanChand & Sons, twentieth edition, 2006.

REFERENCE BOOKS

1. Arun Bahl and B.S. Bahl, Advanced Organic Chemistry, S.Chand & Company Pvt. Ltd., Multicolour edition, 2012.
2. T.W. Graham Solomons, Organic Chemistry, John Wiley & Sons, eleventh edition, 2012.
3. R.T. Morrison and R.N. Boyd, Organic Chemistry, Pearson Education, Asia, sixth edition, 2002.
4. O.P. Agarwal, Organic Chemistry Reactions and Reagents, Goel Publishing house, 2005.

Question paper pattern:

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CORE PRACTICAL – 1
VOLUMETRIC ANALYSIS AND INORGANIC PREPARATION

SUBJECT CODE: 19UCHE302P	PRACTICAL	MARKS 100
SEMESTER: I & II	CREDITS: 3	TOTAL HOURS: 45

Objectives:

To expertise the practical skill in the quantitative chemical analysis and chemical preparations.

A) Volumetric Analysis

1. Estimation of HCl by NaOH using standard oxalic acid solution.
2. Estimation of Na₂CO₃ by HCl using standard Na₂CO₃ solution.
3. Estimation of oxalic acid by KMnO₄ using standard oxalic acid solutions.
4. Estimation of Fe (II) sulphate by KMnO₄ using standard Mohr's salt solution.
5. Estimation of KMnO₄ by thio using standard K₂Cr₂O₇ solution.
6. Estimation of Fe (II) by K₂Cr₂O₇ solution using standard Fe(II) solution.
7. Estimation of Cu(II) sulphate using standard K₂Cr₂O₇ solution.
8. Estimation of Mg(II) by EDTA.
9. Estimation of total hardness of water.
10. Analysis of mixture of carbonate and bicarbonate.
11. Estimation of Vitamin C in fruit juices.

B) Preparation of Inorganic Complexes

1. Ferrous ammonium sulphate
 2. Tetraamminecopper(II) sulphate
 3. Microcosmic salt
 4. Prussian Blue
 5. Hexaamminenickel(II) chloride
-
1. A.L. Vogel, Text book of Inorganic Quantitative Analysis, ELBS, third edition, 1976.
 2. G.S.Vehla, Vogel's Textbook of Macro and Semimicro Qualitative Inorganic Analysis, fifth edition, Revised, 1979.

**NON - MAJOR ELECTIVE
DAIRY CHEMISTRY**

SUBJECT CODE: 19UNME401E	THEORY	MARKS 100
SEMESTER: I	CREDITS: 2	TOTAL HOURS: 30

Objectives:

To learn the chemistry of milk and milk products and also the processes involved in the preservation and formation of milk products.

UNIT - 1: Composition of Milk

Milk - definition - general composition of milk - constituents of milk - lipids, proteins, carbohydrates, vitamins and minerals - physical properties of milk - colour, odour, acidity, specific gravity, viscosity and conductivity - Factors affecting the composition of milk - adulterants, preservatives with neutralizer - examples and their detection - estimation of fat, acidity and total solids in milk.

UNIT - 2: Processing of Milk

Microbiology of milk - destruction of micro - organisms in milk, physico - chemical changes taking place in milk due to processing - boiling, pasteurization - types of pasteurization - Bottle, Batch and HTST (High Temperature Short Time) - Vacuum pasteurization - Ultra High Temperature Pasteurization.

UNIT - 3: Major Milk Products

Cream - definition - composition - chemistry of creaming process - gravitational and centrifugal methods of separation of cream - estimation of fat in cream. Butter - definition - composition - theory of churning – desi butter - salted butter, estimation of acidity and moisture content in butter. Ghee - major constituents - common adulterants added to ghee and their detection - rancidity - definition - prevention - antioxidants and synergists - natural and synthetic.

UNIT - 4: Special Milk

Standardised milk - definition - merits - reconstituted milk - definition - flow diagram of manufacture - Homogenised milk - flavoured milk - vitaminised milk - toned milk - Incitation milk - Vegetable toned milk - humanized milk - condensed milk - definition, composition and nutritive value.

UNIT - 5: Fermented and other Milk Products

Fermented milk products - fermentation of milk - definition, conditions, cultured milk - definition of culture - example, conditions - cultured cream, butter milk - Bulgariious milk - acidophilous milk – Yoheer Indigeneous products - khoa and chhena definition - Ice cream - definition - percentage composition - types - ingredients - manufacture of ice – cream stabilizers - emulsifiers and their role - milk powder - definition - need for making milkpowder - drying process - types of drying.

REFERENCE BOOKS:

1. Robert Jenness and S. Patom, Principles of Dairy Chemistry, S.Wiley, New York, 2005.
2. K. S. Rangappa and K.T. Acharya, Indian Dairy Products, Asia Publishing House NewDelhi, 1974.
3. F.P. Wond, Fundamentals of Dairy Chemistry, Springer, Singapore, 2006.
4. K. Bagavathi Sundari, Applied Chemistry, MJP Publishers, first edition, 2006.
5. Sukumar De, Outlines of Dairy Technology, Oxford University Press, New Delhi, 1980.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 5 out of 10 questions	1 – 10	20	100
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	2	
	Unit – 2	2	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	

**NON-MAJOR ELECTIVE
FOOD CHEMISTRY**

SUBJECT CODE: 19UNME402E	THEORY	MARKS 100
SEMESTER: II	CREDITS: 2	TOTAL HOURS: 30

Objectives:

To create knowledge on chemistry of various food products, adulteration and toxicology.

UNIT - 1: Food Adulteration

Sources of food, types, advantages and disadvantages. Food adulteration - contamination of wheat, rice, milk, butter etc. with clay stones, water and toxic chemicals - Common adulterants, Ghee adulterants and their detection. Detection of adulterated foods by simple analytical techniques.

UNIT - 2: Food Poison

Food poisons - natural poisons (alkaloids - nephrotoxin) - pesticides, (DDT, BHC, Malathion) - Chemical poisons - First aid for poison consumed victims.

UNIT - 3: Food Additives

Food additives - artificial sweeteners - Saccharin - Cyclamate and Aspartate Food flavours - esters, aldehydes and heterocyclic compounds - Food colours - Emulsifying agents - preservatives - leavening agents. Baking powder - yeast - taste makers - MSG - vinegar.

UNIT - 4: Beverages

Beverages - soft drinks - soda - fruit juices - alcoholic beverages - examples. Carbonation - addiction to alcohol - diseases of liver and social problems.

UNIT - 5: Edible Oils

Fats and oils - Sources of oils - production of refined vegetable oils - preservation. Saturated and unsaturated fats - iodine value - role of MUFA and PUFA in preventing heart diseases - determination of iodine value, RM value, saponification values and their significance.

REFERENCEBOOKS:

1. M. Swaminathan, Food Science and Experimental Foods, Ganesh and Company, 1979.
2. Jayashree Ghosh, Fundamental Concepts of Applied Chemistry, S. Chand & Co. Publishers, second edition, 2006.
3. Thanamma Jacob, Text Books of Applied Chemistry for Home Science and Allied Sciences, Macmillan, 1979.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 5 out of 10 questions	1 – 10	20	100
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	2	
	Unit – 2	2	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	

**NON-MAJOR ELECTIVE
FORENSIC CHEMISTRY**

SUBJECT CODE:	THEORY	MARKS 100
SEMESTER: I / II	CREDITS: 2	TOTAL HOURS: 30

Objectives:

To deliver knowledge on chemistry involved in the forensic science.

UNIT - 1: Poisons

Poisons - types and classification - diagnosis of poisons in the living and the dead - clinical symptoms - postmortem appearances. Heavy metal contamination (Hg, Pb, Cd) of sea foods - use of neutron activation analysis in detecting arsenic in human hair. Treatment in cases of poisoning - use of antidotes for common poisons.

UNIT - 2: Crime Detection

Accidental explosion during manufacture of matches and fireworks (as in Sivakasi). Human bombs - possible explosives (gelatin sticks and RDX) - metal detector devices and other security measures for VVIP - composition of bullets and detecting powder burns.

UNIT - 3: Forgery and Counterfeiting

Documents - different types of forged signatures - simulated and traced forgeries - inherent signs of forgery methods - writing deliberately modified - uses of ultraviolet rays - comparison of type written letters – checking silver line water mark in currency notes - alloy analysis using AAS to detect counterfeit coins - detection of gold purity in 22 carat ornaments - detecting gold plated jewels - authenticity of diamond.

UNIT - 4: Tracks and Traces

Tracks and traces - small tracks and police dogs - foot prints - costing of foot prints - residue prints, walking pattern or tyre marks - miscellaneous traces and tracks - glass fracture - tool marks - paints - fibers - Analysis of biological substances - blood, semen, saliva, urine and hair - Cranial analysis (head and teeth) DNA Finger printing for tissue identification in dismembered bodies - detecting steroid consumption in athletes and race horses.

UNIT - 5: Medical Aspects

Aids - causes and prevention - misuse of scheduled drugs - burns and their treatment by plastic surgery. Metabolite analysis using mass spectrum - Gas chromatography-Arson - natural fires and arson - burning characteristics and chemistry of combustible materials - nature of combustion. Ballistics - classification - internal and terminal ballistics - small arms - laboratory examination of barrel washing and detection of powder residue by chemical tests.

REFERENCEBOOKS

1. T. H. James, Forensic Sciences, Stanley Thornes Ltd, 1987.
2. Richard Saferstin and Criminalistics - An Introduction to Forensic Science (College Version), Sopfestein, Printice hall, eighth edition, 2003.

**NON-MAJOR ELECTIVE
CHEMISTRY IN EVERY DAY LIFE**

SUBJECT CODE:	THEORY	MARKS 100
SEMESTER: II	CREDITS: 2	TOTAL HOURS: 30

Objectives:

To learn chemistry involved in our day-by-day life.

UNIT –1

General survey of chemicals used in everyday life. Air - components and their importance; photosynthetic reaction, air pollution, green - house effect and the impact on our life style. Water - Sources of water, qualities of potable water, soft and hard water, methods of removal of hardness - water pollution.

UNIT –2

Building materials - cement, ceramics, glass and refractories - definition, composition and application only. Plastics - polythene, PVC, bakelite, polyesters, melamine- formaldehyde resins - preparation and uses only.

UNIT – 3

Food and Nutrition - Carbohydrates, Proteins, Fats - definition and their importance as food constituents - balanced diet - Calorie - minerals and vitamins (sources and their physiological importance). Cosmetics - tooth paste, face powder, soaps and detergents, shampoos, nail polish, perfumes - general formulation and preparations - possible hazards of cosmetic use.

UNIT – 4

Chemicals in food production - fertilizers - need, natural sources; urea, NPK fertilizers and super phosphate. Fuel - classification - solid, liquid and gaseous; nuclear fuel examples and uses.

UNIT – 5

Pharmaceutical drugs - analgesics and antipyretics - paracetamol and aspirin. Colour chemicals - pigments and dyes - examples and applications. Explosives - classification and examples.

REFERENCE BOOKS:

1. Randolph. Norris Shreve, Chemical Process Industries, McGraw-Hill, Texas, fourth edition, 1977.
2. W. A. Poucher, Joseph A. Brink, Jr. Perfumes, Cosmetics and Soaps, Springer, 2000.
3. A. K. De, Environmental Chemistry, New Age International Public Co., 1990.

Question paper pattern:

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 5 out of 10 questions	1 – 10	20	100
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	2	
	Unit – 2	2	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	

SEMESTER - III

CORE THEORY – 3

ORGANO-OXYGEN COMPOUNDS

SUBJECT CODE: 19UCHE304	THEORY	MARKS 100
SEMESTER: III	CREDITS: 4	TOTAL HOURS: 60

UNIT-I : CHEMISTRY OF ALCOHOLS, ETHERS AND PHENOLS. (15 HRS)

Monohydric alcohols: Nomenclature, preparation of alcohols from alkenes (by oxymercuration -demercuration), alkyl halides, Grignard reagent and carbonyl compounds (by reduction). Reactions of alcohols- Dehydration, oxidation, action of hot reduced copper.

Ethers: Nomenclature, preparation by Williamson ether synthesis. Reactions-acid catalysed cleavage of ethers. Epoxides (Oxirane) - preparation and reactions of epoxides.

Phenols: Nomenclature, synthesis of phenol from benzene sulphonic acid, chlorobenzene and cumene. Properties – Acidity of phenols (explanation on the basis of resonance stabilization). Electrophilic substitution reaction of phenol (orientation of electrophiles) - halogenation, nitration, sulphonation, Reimer-Tiemann reaction, Kolbe-Schmidt reaction and coupling with diazonium salts.

UNIT-II : CHEMISTRY OF CARBONYL COMPOUNDS (10 HRS)

Nomenclature, structure of carbonyl compounds, acidity of alpha-hydrogen atom. Mechanism of nucleophilic addition with RMgX , HCN , ROH , NaHSO_3 , NH_3 and its derivatives.

Mechanism of Meerwein-Ponndorf-Verley reduction, Clemmensen reduction, Wolf-Kishner reduction, aldol condensation, Claisen-Schmidt reaction, Cannizzaro reaction, haloform reaction, Knoevenagel, Perkin and Benzoin condensation reactions.

UNIT-III: CHEMISTRY OF CARBOXYLIC ACIDS AND THEIR DERIVATIVES.

(15 HRS)

Acidity of carboxylic acids, effect of substituents on acidity, comparison of acid strengths of halogen substituted acetic acids, acid strengths of substituted benzoic acids.

Dicarboxylic acids: Preparation – from alkyl cyanides, cyclic ketones and haloesters. Reactions – action of heat, action of PCl_5 and NH_3 .

Acid derivatives (Aliphatic): Synthesis and important properties of acid derivatives (acid chlorides, acid anhydrides, esters and amides). Relative reactivity of acyl compounds.

UNIT –IV: ACTIVE METHYLENE COMPOUNDS

(10 HRS)

Active methylene compounds - keto-enol tautomerism. Compounds containing active methylene group- acetoacetic, malonic and cyanoacetic esters. Preparation and synthetic applications.

UNIT –V: CARBOHYDRATES

(10 HRS)

Carbohydrates – Definition and classification of carbohydrates with examples. Mono saccharides: Explanation of enantiomers, diastereomers, epimers and anomers with examples. Mechanism of mutarotation, Osazone formation.

Absolute configurations of glucose and fructose (includes cyclic and Haworth structure). Interconversion (aldose to ketose, ketose to aldose, arabinose to glucose, glucose to arabinose and glucose to mannose). Disaccharides – Sucrose, Maltose – structural elucidation. Polysaccharides -Starch and cellulose (Elementary treatments)

TEXT BOOKS

1. M. K. Jain, S. C. Sharma, Modern Organic Chemistry, Vishal Publishing, fourth reprint, 2003.
2. S.M. Mukherji, and S.P. Singh, Reaction Mechanism in Organic Chemistry, MacmillanIndia Ltd., third edition, 1994.
3. Arun Bahl and B.S. Bahl, Advanced organic chemistry, New Delhi, S.Chand & Company Pvt. Ltd., Multicolour edition, 2012.
4. R. T. Morrison, R. N. Boyd, and S. N. Bhattacharjee, Organic Chemistry, Pearson Education, Asia, seventh edition, 2012.

REFERENCE BOOKS

1. T. W. Graham Solomons, Organic Chemistry, John Wiley & Sons, fifth edition, 1992.
2. A. Carey Francis Organic Chemistry, Tata McGraw-Hill Education Pvt., Ltd., New Delhi, seventh edition, 2009.
3. I. L. Finar, Organic Chemistry, Wesley Longman Ltd, England, sixth edition, 1996.
4. P. L. Soni, and H. M. Chawla - Text Book of Organic Chemistry, New Delhi, Sultan Chand & Sons, twenty ninth edition, 2007.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

SEMESTER - IV

CORE THEORY – 4
CHEMISTRY OF s-BLOCK & p-BLOCK ELEMENTS

SUBJECT CODE: 19UCHE306	THEORY	MARKS 100
SEMESTER: IV	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To know the nature of compounds formed by s - and p - block elements; to understand the aspects in gaseous, liquid and solid states; to understand the crystal structures of ionic compounds.

UNIT - 1: Chemistry of s-Block Elements **(10 hrs)**

Characteristic properties of s-block elements, preparation, properties and uses of NaOH, Na₂CO₃, KBr and KClO₃. Hydrides – Classification as ionic, molecular and metallic hydrides, preparation, important properties and uses of LiAlH₄ and CaH₂. Comparative study of the element with respect to oxides, hydroxides, halides, carbonates and bicarbonates. Diagonal relationship of Li and Mg. Anomalous behavior of Li and Be. Extraction of beryllium. Chemical properties of metals: reaction with water, air and nitrogen.

UNIT - 2: Boron and Carbon Family **(15 hrs)**

Extraction of B and Si - Compounds of boron with oxygen - preparation, properties and uses of boron sesquioxide, borates, borax, sodium peroxoborates. - Compounds of Boron with nitrogen - preparation, properties and uses of boron nitride, borazine. Boron hydrides - preparation, properties, uses and structure of diborane, reaction with ammonia, hydroboration.

Aluminium - Extraction of Aluminium and its uses, Alloys of Aluminium, amphoteric behavior, aluminates. Comparison of carbon with silicon, Manufacture and uses of CO and CO₂ (includes dry ice); allotropy of carbon and carbides. Preparation, types and industrial applications of carbides, interstitial carbides and covalent carbides.

Silicates - Types of silicates – application of silicates in technology - alkali silicates, ceramics and glass. Preparation, properties and uses of silicones. Tin –

Allotropic forms of Tin, alloys of tin, tinning, tin plating. Lead - lead accumulator (discharging and recharging), lead pigments.

UNIT - 3: Nitrogen and Oxygen Group Elements (15 hrs)

Metallic and non metallic character of group 15 elements; hydrides and halides of group-15 elements - hydrazine, hydroxylamine, phosphene, ammonium nitrate, sodium bismuthate - properties and uses. Oxides of group 15 elements: oxides of nitrogen - dinitrogen tetroxide, dinitrogen pentoxide; oxides of phosphorus - oxoacids of nitrogen: nitrous acid, nitric acid, hyponitrous acid; oxoacids of phosphorus - orthophosphorous acid, metaphosphorous acid, hypophosphorous acid; orthophosphoric acid, di-, tri- and tetra polyphosphoric acids;

Group-16 (oxygen group): Ozone, oxides - normal oxides, peroxides, suboxides, basic oxides, amphoteric oxides, acidic oxides, neutral oxides. Oxides of sulphur - SO_2 , SO_3 ; oxoacids of sulphur - thionic acid series, peroxyacid series, oxohalides - thionyl compounds, sulfuryl compounds (methods of preparation and properties).

UNIT - 4: Halogens (10 hrs)

Ionic – covalent - bridging halides, reactivity of halogens, reduction of halogens by thiosulphate. Halogen oxides: oxygen difluoride, dioxygen difluoride, dichlorine monoxide, chlorine dioxide, dichlorine hexoxide, dichlorine heptoxide; bleaching powder - estimation of available chlorine; bromine dioxide, iodine pentoxide. Oxoacids of halogens: hypohalous acid (HOX), halous acid (HXO_2), halic oxide (HXO_3), perhalic acid (HXO_4), strength of oxoacids. Inter-halogen compounds: ClF , ICl , ClF_3 , BrF_3 , IF_3 , ClF_5 , BrF_5 , IF_5 ; poly halides.

UNIT - 5: Noble Gases (10 hrs)

Noble gases: position in the periodic table – general characteristics – structure and shape of Xenon compounds – XeF_2 , XeF_4 , XeF_6 , XeOF_4 – uses of Noble gases. Chemical reactivity of noble gases, preparation, structure and bonding of noble gas compounds. Chemistry of xenon: structure and bonding of

xenon fluorides, - oxides and oxyfluorides of xenon.

TEXT BOOKS

1. B. R. Puri, L. R. Sharma and K. C. Kalia, Principles of Inorganic Chemistry, Milestonepublishers and distributors, thirteenth edition,2009.
2. P. L. Soni and Mohan katyal, Textbook of Inorganic Chemistry, Sultan Chand & Sons,twentieth edition,2006.
3. J. D. Lee, Concise Inorganic Chemistry, Blackwell science, fifth edition, 2005.

REFERENCE BOOKS

1. A. K. De, Text book of Inorganic Chemistry, Wiley East Ltd, seventh edition, 1992.
2. W.U. Malik, G.D. Tuli and R.D. Madan, Selected Topics in Inorganic Chemistry; S. Chand and Co, 2014.
3. R. C. Agrawal, Modern Inorganic Chemistry, Kitab Mahal, first edition, 1987.
4. K. N. Upadhayaya, Text book of Inorganic Chemistry, Vikas Publishing House, NewDelhi, 1990.
5. Peter Atkins, Tina Overton, Jonathan Rourke and Mark Weller, Inorganic Chemistry,Oxford University Press, sixth edition, 2014.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CORE PRACTICAL – 2
INORGANIC QUALITATIVE ANALYSIS

SUBJECT CODE: 19UCHE305P	PRACTICAL	MARKS 100
SEMESTER: III & IV	CREDITS: 3	TOTAL HOURS: 45

Objectives:

To develop the skill on systematic analysis of inorganic salts.

Semi - Micro Qualitative Analysis

1. Analysis of simple acid radicals:
Carbonate, sulphide, sulphate, thiosulphite, chloride, bromide, iodide, nitrate
2. Analysis of interfering acid radicals:
Fluoride, oxalate, borate, phosphate, arsenate, arsenite.
3. Elimination of interfering acid radicals and Identifying the group of basic radicals
4. Analysis of basic radicals (group wise):
Lead, copper, bismuth, cadmium, tin, antimony, iron, aluminium, arsenic, zinc, manganese, nickel, cobalt, calcium, strontium, barium, magnesium, ammonium
5. Analysis of a mixture - I to VIII containing two cations and two anions (of which one is interfering type)

REFERENCE BOOK

V. Venkateswaran, R. Veeraswamy and A. R. Kulandivelu, Basic Principles of Practical Chemistry, Sultan Chand & Sons, New Delhi, second edition, 1997.

CORE THEORY – 5
ORGANO NITROGEN COMPOUNDS AND NATURAL PRODUCTS

SUBJECT CODE: 19UCHE307	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To introduce chemistry of nitrogen compounds and natural products namely aminoacids, proteins and alkaloids.

UNIT – 1: Amines, Diazonium Compounds and Nitro Compounds (15 hrs)

Aliphatic and aromatic amines: Preparation of primary, secondary and tertiary amines. Reactions: basicity of amines, effect of substituent on basicity of aromatic amines. Some sulpha drugs, Diazonium salts: Preparation, diazotization reactions, replacement reactions (Sandmeyer, Gatterman and Gomberg reactions), coupling reactions.

Nitro compounds: Nomenclature and classification, aliphatic and aromatic nitro compounds, general properties, preparation by nitration. Reactions: reduction by chemical and electrolytic methods. Di- and tri- substitution of aromatic nitro compounds: synthesis of o-, m-, p- dinitrobenzenes and trinitrobenzene.

UNIT –2: Amino acids and Proteins (15 hrs)

Classification, structure and stereochemistry of amino acids, Acid - base behavior isoelectric point and electrophoresis, preparation and reactions of α - amino acids, structure and nomenclature of peptides and proteins, classification of proteins, peptide structure determination, end group analysis, selective hydrolysis of peptides, classical peptide synthesis, solid - phase peptide synthesis, Structure of peptides and proteins, Protein denaturation / renaturation.

UNIT –3: Nucleic acids and Heterocyclic Compounds (10 hrs)

Introduction, constituents of nucleic acids, nucleosides and nucleotides. The double helical structure of DNA. Molecular orbital picture and aromatic

characteristics of pyrrole, furan, thiophene and pyridine. Methods of synthesis and chemical reactions with particular emphasis on the mechanism of electrophilic substitution, mechanism of nucleophilic substitution reaction in pyridine derivatives, comparison of basicity of pyridine, piperidine and pyrrole.

Introduction to condensed five and six membered heterocycles. Preparation and reactions of indole, quinoline and isoquinoline with special reference to Fisher indole synthesis, Skraup synthesis, Bischler-Napieralski synthesis. Electrophilic substitution reactions mechanism of indole, quinoline and isoquinoline.

UNIT –4: Alkaloids

(10 hrs)

Alkaloids: Definition, Occurrence, extraction of alkaloids from plants, general properties, determination of the chemical constitution of the alkaloids, functional group analysis, estimation of groups, degradation and synthesis. Structural elucidation of Coniine, Piperine and Nicotine.

UNIT –5: Terpenoids

(10 hrs)

Terpenoids: Classification, isoprene rule, isolation and general properties, occurrence, structural elucidation of geraniol, citral, menthol, α - pinene and camphor.

REFERENCE BOOKS

1. R. T. Morrison and R. N. Boyd, Organic Chemistry, Pearson Education, Asia, sixth edition, 2002.
2. T.W. Graham Solomons, Organic Chemistry, John Wiley & Sons, eleventh edition, 2012.
3. A. Carey Francis, Organic Chemistry, Tata McGraw-Hill Education Pvt. Ltd., New Delhi, seventh edition, 2009.
4. I. L. Finar, Organic Chemistry, Vol. (1 & 2), England, Wesley Longman Ltd, sixth edition, 1996.
5. J. A. Joule, and G. F. Smith, Heterocyclic Chemistry, Van Nostrand Reinhold (UK) Co.Ltd., England, second edition, 1984.
6. P. L.Soni and H. M. Chawla, Text Book of Organic Chemistry, Sultan

Chand & Sons, New Delhi, twenty ninth edition, 2007.

Question paper pattern:

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

SEMESTER - V

CORE THEORY – 6
CHEMICAL KINETICS AND ELECTROCHEMISTRY

SUBJECT CODE: 19UCHE308	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To understand the concepts of kinetics of chemical reaction and electrochemistry.

UNIT –1: Chemical Kinetics – I **(10 hrs)**

Rate of reaction - average and instantaneous rate - rate equation – order and molecularity - rate laws - rate constants - derivation of rate constants and characteristics of first, second, third and zero order reactions - derivation of time for half change with examples methods for determination of order of reaction – experimental methods of determination of rate constant of reactions – volumetry, manometry and polarimetry.

UNIT –2: Chemical Kinetics – II **(15 hrs)**

Effect of temperature on reaction rates - concept of activation energy - Arrhenius equation - Collision Theory - derivation of rate constant for bimolecular reactions - failure of collision theory - Lindemann's Theory of unimolecular reactions – theory of absolute reaction rates - derivation of rate constant for bimolecular reaction - significance of entropy and free energy of activation - comparison of collision theory and ARRT - consecutive, parallel and reversible reactions (no derivation, only examples)

Catalysis: Definition - characteristics of a catalyst - homogeneous and heterogeneous catalysis - function of a catalyst in terms of Gibb's free energy of activation - kinetics of acid-base and enzyme catalysis - heterogeneous catalysis - kinetics of unimolecular surface reactions.

Adsorption: Physisorption and chemisorption - Freundlich adsorption isotherm - Langmuir adsorption isotherm – applications of adsorption.

UNIT - 3: Electrical Conductance

(15 hrs)

Electrical transport, conductance in metal and in electrolytic solution. Specific conductance and equivalent conductance. Measurement of equivalent conductance using Kohlraush's bridge - Arrhenius theory of electrolytic dissociation and its limitation - Weak and strong electrolyte according to Arrhenius theory, Ostwald's dilution law - applications and limitation - Variation of equivalent conductance with concentration.

Migration of ion, ionic mobility. Kohlraush's law and its applications. The elementary treatment of the Debye-Huckel-Onsager equation for strong electrolytes, evidence for ionic atmosphere, the conductance at high fields (Wien effect) and high frequencies (Debye - Falkenhagen effect), Transport number and Hittorf rule. Determination by Hittorf method and moving boundary method. Application of conductance measurements - Determination of λ_0 of strong electrolytes. Determination of K_a of weak acids. Determination of solubility product of a sparingly soluble salt, common ion effect and conductometric titrations.

UNIT - 4: Electrochemical Cells – I

(10 hrs)

Electrolytic and Galvanic cells – Reversible and irreversible cells. Conventional representation of electrochemical cells. Electromotive force of a cell and its measurement - calculation of thermodynamic quantities of cell reactions (ΔG , ΔH , ΔS and K) - application of Gibbs-Helmholtz equation and Nernst equation.

Types of reversible electrodes – Gas/metal - metal ion/ metal - insoluble salt and redox electrodes. Electrode reactions – Nernst equation – Derivation of cell EMF and single electrode potential, standard hydrogen electrode – reference electrodes – standard electrode potentials – sign convention – Electrochemical series and its significance.

UNIT - 5: Electrochemical Cells – II

(10 hrs)

Concentration cell with and without transport - Liquid junction potential - Application of EMF concentration cells. Valence of ion, solubility product and

activity co-efficient. Potentiometric titrations. Determination of pH using hydrogen, quinhydrone and glass electrodes - Corrosion – general and electrochemical theory – passivity – prevention of corrosion.

TEXT BOOKS

1. B.R. Puri, L.R. Sharma and M.S. Pathania, Principles of Physical Chemistry, ShobanlalNagin Chand and Co. Jalendhar, forty first, edition, 2001
2. S. H. Maron and J. B. Lando, Fundamentals of Physical Chemistry, Macmillan limited, New York, 1992.
3. S. K. Dogra and S. Dogra, Physical Chemistry through Problems: New Age International, fourth edition, 1996.

REFERENCE BOOKS

1. Gilbert. W. Castellen, Physical Chemistry, Narosa Publishing House, third edition, 1985.
2. P. W. Atkins, and Julio de Paula, Physical Chemistry, Oxford University press, seventh edition, 2002.
3. K. L. Kapoor, A Textbook of Physical Chemistry, Macmillan India Ltd, third edition, 2009.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CORE THEORY – 7
SPECIAL TOPICS IN CHEMISTRY

SUBJECT CODE: 19UCHE309	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To emphasize the fundamentals of stereochemistry, group theory and photochemistry, molecular rearrangement and their applications.

UNIT - 1: Stereochemistry – I (10 hrs)

Stereoisomerism – classification; Optical isomerism – optical activity, specific rotation (calculation) – asymmetric center – chirality – achiral molecules. Elements of symmetry. Optical activity of biphenyls, allenes and spiranes. Racemisation – methods of racemisation (by substitution and tautomerism); resolution and methods of resolution (mechanical, biochemical and chemical). Asymmetric synthesis: Partial and absolute – Cram's rule (Elementary approach only).

UNIT –4: Stereochemistry – II (15 hrs)

Projection formula – flying wedge, Fischer, Sawhorse and Newmann – interconversion - d, l and D, L notations of optical isomers – CIP rules – R and S nomenclature - optical isomers up to three chiral carbon atoms. Meso, dl, erythro and threo representations.

Geometrical isomerism – cis and trans, syn-anti and E-Z nomenclature. Geometrical isomerism in maleic, fumaric acids and unsymmetrical ketoximes. Methods of distinguishing geometrical isomers. Geometrical isomerism of disubstituted cyclohexanones.

Conformational analysis - conformers, configuration, dihedral angle and torsional strains. Conformational analysis of ethane and disubstituted ethane derivatives, conformers of cyclohexane, bonds - ring flipping, mono and disubstituted cyclohexanes.

UNIT –3: Group Theory (15 hrs)

Symmetry elements and symmetry operations - symmetry operation of water molecule, illustration of mathematical rules for the group using symmetry

operations of water molecule. Construction of group multiplication table for water molecule. Point groups - definition. Symmetry elements and symmetry operations of the following groups – C_2 , C_3 , C_{2v} , C_{3v} , C_{2h} .

UNIT –4: Photochemistry

(10 hrs)

Laws of photochemistry – Lambert - Beer's law – Grothus – Draper law and Stark Einstein law–Quantum efficiency –reasons for high and low quantum yield - comparison between thermal and photochemical reactions – rate law – kinetics of H_2 and Cl_2 ; H_2 and Br_2 reactions.

UNIT –5: Molecular Rearrangements

(10 hrs)

Types of rearrangements - mechanism of pinacol-pinacolone, benzyl-benzilic acid, bezindine, Favorski, dienone- phenol, Claisen, Fries, Hoffmann, Curtius, Schmidt and Beckmann rearrangements.

TEXT BOOKS

1. B. R. Puri, Sharma. L. R and Kalia, K. C, Principles of Inorganic Chemistry, S. Chand Publications, New Delhi, twenty third edition, 1993.
2. J. D. Lee, Concise Inorganic Chemistry, UK Blackwell Science, (Oxford), fifth edition, 1996.
3. P.L. Soni, Mohankatyal, Text Book of Inorganic Chemistry, Sultan Chand & sons, New Delhi, twentieth edition, 2006.
4. S. Swarnalakshmi, T. Saroja, R. M. Ezhilarasi; A Simple Approach to Group Theory in Chemistry Universities Press (India) Pvt. Ltd., first edition, 2008.
5. M. K. Jain, S. C. Sharma Modern Organic Chemistry, Vishal Publishers, fourth edition, 2009.
6. B.R. Puri Sharma and M.S. Pathania, Principles of Physical Chemistry, Vishal Publishing Co., forty second, edition, 2008.
7. P. S. Kalsi, Stereochemistry - Conformation & Mechanism, New Age International limited, sixth edition, 2006.
8. P. S. Kalsi, Stereochemistry and Mechanisms through Solved Problems, New Age International limited, third edition, 2001.

REFERENCE BOOKS

1. S.H. Maron and C.F. Prutton, Physical Chemistry, McMillan, London fourth edition, 1965.
2. P. W. Atkins and Julio de Paula, Physical Chemistry, Oxford University Press, seventh edition, 2002.
3. G. W. Castellan, Physical Chemistry, Narosa Publishers, fourth edition, 2004.
4. H. J. Arnikar, Essentials of Nuclear Chemistry, New Age International Publishers, fourth edition, 1995.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
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TOTAL MARKS				100

Distribution of Questions:

Sections	Units	No. of Questions	
		Theory	Problems
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	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CORE THEORY – 8
ANALYTICAL CHEMISTRY-I

SUBJECT CODE: 19UCHE310	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives:

- *To build a basic knowledge on generation of analytical data in an appropriate manner; to provide thorough knowledge on gravimetry.*
- *To expertise the instrumental methods of chemical analysis for microgram level; to cultivate the analytical skill in the structural identification of chemical compounds.*

UNIT –1: Data Analysis and Sampling (15 hrs)

Precision, accuracy, theory of errors, idea of significant figures and its importance with examples, methods of expressing accuracy, error analysis, minimizing errors, methods of expressing precision, average, deviation, standard deviation and confidence limit; analytical balance; elementary aspect of calibration, calibration of glassware and other equipments for quantitative chemical analysis.

Sampling – Significance of sampling, types of sample, sampling methods for solids, liquids and gases.

UNIT –2: Gravimetry (10 hrs)

Gravimetric analysis – principle, precipitating agent, condition of precipitation; selective and specific precipitants–DMG, cupferron, salicylaldehyde, ethylenediamine; use of sequestering agent, coprecipitation and post precipitation, peptisation, calculations in gravimetric method and gravimetric factor.

UNIT –3: Separation Techniques (15 hrs)

Separation by precipitation; solvent extraction–types and applications; Chromatographic techniques – types, principle, theory and applications of thin

layer, paper chromatography, column and ion-exchange chromatography; gas liquid chromatography and high performance liquid chromatography.

UNIT –4: Purification Techniques

(10 hrs)

Drying of solid, distillation-principle, types, and experiments of simple distillation, fractional, steam, azeotrope and vacuum distillation; Soxhlet extraction, recrystallization, fractional crystallization, sublimation. Testing of purity – determination of melting point, boiling point, refractive index and density.

UNIT –5: Polarimetry and Thermal analysis

(10 hrs)

Polarimetry – principle, instrumentation and applications; estimation of glucose. Thermo-analytical methods: Principle involved in thermo gravimetric analysis and differential thermal analysis, discussion of various components with block diagram, characteristics of TGA and DTA – factors affecting TGA and DTA curves.

TEXT BOOKS

1. R. Gopalan, K. Rengarajan, P. S. Subramanian, Elements of Analytical Chemistry, Sultan Chand & Sons, third edition, 2003.
2. David Harvey, Modern Analytical Chemistry, Tata McGraw-Hill, first edition, 2000.
3. J. Mendham, R. C. Denney, J. D. Barnes and M. Thomas, Vogel's Text book of Quantitative Chemical Analysis, Pearson Education Pvt. Ltd, sixth edition, 2004.

REFERENCE BOOKS

1. Douglas A. Skoog, Donald M. West and F. James Holler, Fundamentals of Analytical Chemistry Harcourt Asia Pvt. Ltd., ninth edition, 2001.
2. Douglas A. Skoog, Donald M. West and F. James Holler, Analytical Chemistry, An Introduction; Saunders College Publishers, seventh edition, 2000.
3. Dean, John A. Merritt, Lynne L., Jr. Settle, Frank A., Jr. Willard, Hobart H,

Instrumental Methods of Analysis, Wadsworth Publishing Co Inc., seventh edition, 1988.

4. Skoog D. A., Principles of Instrumental Analysis, Saunders College Publishing, fifth edition, 1998.

Visits to

- i. Quality Control laboratories.
- ii. Various chemical industries with established laboratories.
- iii. BIS & FSSAI approved establishments.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

SEMESTER - VI

CORE THEORY-9
THERMODYNAMICS AND SOLUTIONS

SUBJECT CODE: 19UCHE314	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives:

To strengthen the knowledge on Thermodynamics and Solutions.

UNIT – 1: Thermodynamics I (15 hrs)

Definition and explanation of terms - System –surrounding - open, closed and isolated systems - state of system - intensive and extensive properties - Thermodynamic equilibrium - Thermodynamic processes - isothermal, adiabatic, reversible, irreversible and cyclic processes - state and path functions - exact and inexact differentials - concept of heat and work - work of expansion at constant pressure and free expansion.

First law of Thermodynamics - statement and equation - definition of internal energy(U), Enthalpy(H) and heat capacity - relation between C_p and C_v - Calculation of W, q, ΔU and ΔH for expansion of ideal gases under reversible isothermal and adiabatic conditions

– Joule - Thomson effect - Calculation of μ_{JT} for ideal and real gases - inversion temperature and its significance. Thermo chemistry - Relation between enthalpy of reaction at constant volume (q_v) and at constant pressure(q_p), variation of enthalpy of reaction with temperature - Kirchhoff's equation - enthalpy of combustion - flame and explosion temperature - heat of neutralization - heat of formation - integral heat of solution and dilution - Bond energy and its calculation from thermo chemical data - Hess's law and its applications.

UNIT –2: Thermodynamics II (10 hrs)

Second law of Thermodynamics - need for second law - different statements of second law - spontaneous process - concept of entropy – definition - entropy of an ideal gas - entropy changes in reversible, cyclic and physical

transformations - physical significance of entropy - calculation of entropy changes with changes in P, V, T and entropy of mixing - Entropy criterion for spontaneous and equilibrium processes - Gibb's free energy(G) and Helmholtz free energy (A) – variation of G and A with P, V and T - Criteria for Spontaneity - Gibb's - Helmholtz equation and its applications.

UNIT –3: Thermodynamics III

(10 hrs)

Third law of Thermodynamics - Nernst heat theorem - statement of third law of Thermodynamics - determination of absolute entropy from heat capacity measurements - exceptions to third law.

Thermodynamic treatment of law of mass action – van't Hoff reaction isotherm - Standard free energy change with equilibrium constant - Variation of equilibrium constant with temperature change - van't Hoff isotherm. Partial molar properties – Basic concept of Chemical potential - fugacity, activity and activity coefficient– Clapeyron-Clausius equation -derivation and its uses.

UNIT –4: Solutions and Colloids

(15 hrs)

Ideal and non - ideal solutions - solutions of liquids in liquids - Raoult's law - binary liquid mixtures - deviations from ideal behaviour - vapour pressure – composition and vapour pressure - temperature curves - azeotropic distillation – Ethanol-water system and HCl-Water system - partially miscible liquid systems - phenol - water system, triethylamine - water, Nicotine - water system - effect of impurities on CST - completely immiscible liquids - steam distillation - solutions of gases in liquids - Henry's law.

Dilute solutions: Colligative properties - relative lowering of vapour pressure - thermodynamic derivation of elevation of boiling point - depression of freezing point – calculation of molecular weight – osmosis - laws of osmotic pressure- distribution law - thermodynamic derivation and application.

UNIT –5: Phase Equilibria

(10 hrs)

Definition of terms in phase rule – derivation and application to one component system - water and sulphur – super cooling, sublimation. Two

component systems – Solid liquid equilibria, simple eutectic lead – silver (desilverisation of lead), Bi-Cd; compound formation with congruent melting point (Mg - Zn, FeCl₃ – H₂O) and incongruent melting point (Na - K, Na₂SO₄ - H₂O).

REFERENCE BOOKS

1. B.R. Puri and L.R. Sharma, Principles of Physical Chemistry, ShobanLalNagin Chandand Co., thirty three edition, 1992.
2. S.H. Maron and J.B. Lando, Fundamentals of Physical Chemistry, Macmillan limited, New York, 1966.
3. S.K. Dogra and S. Dogra, Physical Chemistry Through Problems, New age international, fourth edition, 1996.
4. Gilbert. W. Castellan, Physical Chemistry, Narosa publishing house, third edition, 1985.
5. P. W. Atkins, and Julio de Paula, Physical Chemistry, Oxford University press, seventh edition, 2002.
6. J. Rajaram and J.C. Kuriacose, Thermodynamics, ShobanLalNagin Chand and CO., 1986.
7. K. L. Kapoor, A Textbook of Physical chemistry, (volume-2 and 3), Macmillan, India Ltd, third edition, 2009.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
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TOTAL MARKS				100

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	Unit – 3	1	
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	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit – 5	1	

CORE THEORY-10
CHEMISTRY OF d-BLOCK & f-BLOCK ELEMENTS AND COORDINATION
CHEMISTRY

SUBJECT CODE: 19UCHE315	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives

To learn the chemistry of elements present in the d- and f-blocks of periodic table.

UNIT - 1: Transition Elements **(15 hrs)**

Characteristic properties of d-block elements: magnetic property, reactivity, variable oxidation states, catalytic properties, comparative study of the elements of the first transition series with reference to size, ionization potentials, redox potentials, magnetic behaviour, oxidation states and ability to form complex compounds, trends in chemical and physical properties in passing from first to second and third series; important uses of transition metals and their alloys. Extraction of Ti, V, W, Cr.

UNIT - 2: Chemistry of f-Block Elements **(10 hrs)**

General characteristics of f-block elements- Lanthanides: lanthanide contraction, similarity in properties, occurrence, oxidation states, chemical properties of Ln(III) cations, magnetic properties. Separation of lanthanides: solvent extraction, ion exchange.

Actinides: occurrence, oxidation states, magnetic properties. Extraction of Uranium and thorium. Preparation, properties and uses of ceric ammonium sulphate, thorium dioxide, thorium nitrate, uranium hexafluoride, uranyl acetate. Comparison of lanthanides and actinides.

UNIT - 3: Coordination Chemistry **(15 hrs)**

Review of the fundamentals of coordination chemistry: Distinction between double salts and coordination compounds - Terminology, types of ligands - monodentate, bidentate, polydentate and ambidentate ligands, IUPAC rules for nomenclature of coordination compounds, applications of complexes in

qualitative and quantitative analysis.

Isomerism: linkage, ionization, hydrate, coordination, coordination position isomerism, geometrical (cis and trans; facial and meridional), optical isomerism in 4 - coordinated and 6-coordinated complexes, trans effect, stability of coordination compounds – overall and step wise stability constants. Factors affecting the stability of complexes.

UNIT - 4: Theories of Coordination Complexes (10 hrs)

Werner's coordination theory, Sidgwick theory - EAN rule and stability, valence bond theory, 4-coordinate complexes, 6-coordinate complexes - Inner and outer orbital complexes, hybridization, geometry, magnetism, limitations of VBT. Crystal field theory: crystal field effects, crystal field splitting in octahedral, tetragonally distorted octahedral geometry, tetrahedral geometries and square planar complexes - high spin and low spin complexes; CFSE and factors affecting it; computation of CFSE; evidences of crystal field splitting; spectrochemical series.

UNIT - 5: Reaction Mechanisms, synthesis and biological importance of Coordination Compounds (10 hrs)

Substitution reactions in octahedral complexes and square planar complexes: dissociative and associative and interchange mechanisms (simple approach). Cis- and trans- effects in synthesis of square planar and octahedral complexes. Biological importance of transition metals: biological roles of Cr, Mo, Mn, Fe, Co, Cu, Zn.

TEXTBOOKS

1. B.R. Puri, L.R. Sharma and K.C. Kalia, Principles of Inorganic Chemistry, Milestone publishers, New Delhi, thirtieth edition, 2009.
2. P.L. Soni and Mohan kalyal, Textbook of Inorganic Chemistry, Sultan Chand & Sons, twentieth edition, 2006.
3. J.D. Lee, Concise Inorganic Chemistry, Blackwell science, fifth edition, 2005.

REFERENCE BOOKS

1. A. K. De, Text book of Inorganic Chemistry, Wiley East, seventh edition, 1992.
2. Wahid U Malik, G. D.Tuli and R. D. Madan, Selected Topics in Inorganic Chemistry S. Chand and Co, nineteenth edition, 2014.
3. R. C. Agrawal, Modern Inorganic Chemistry, Kitab Mahal, first edition, 1987.
4. K. N. Upadhyaya, Text book of Inorganic Chemistry, Vikas Publishing House, NewDelhi,1990.

Question paper pattern

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	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CORE THEORY – 11
ANALYTICAL CHEMISTRY–II

SUBJECT CODE: 19UCHE316	THEORY	MARKS 100
SEMESTER: V	CREDITS: 4	TOTAL HOURS: 60

Objectives

To provide a knowledge on instrumental methods of analytical techniques and spectroscopy.

UNIT –1: Polarography and Diffraction Methods (15 hrs)

Polarography - principle, concentration polarization, dropping mercury electrode, advantages and disadvantages, migration and diffusion currents. Ilkovic equation and significance. Instrumentation - experimental assembly and electrodes. Current voltage curve, oxygen wave - influence of temperature and agitation on diffusion layer. Applications - qualitative and quantitative applications for inorganic systems.

Theory, instrumentation and applications of X - ray, electron and neutron diffraction analyses.

UNIT –2: Atomic and Molecular Spectroscopy (15 hrs)

Principle, instrumentation and applications of Atomic Absorption Spectroscopy and Flame Photometry.

Ultraviolet and Visible spectroscopy: theory, instrumentation and application (qualitative and quantitative) – photometric titrations. Infrared Spectroscopy: Theory, instrumentation and applications. Raman Spectroscopy: Theory, instrumentation and application.

UNIT –3: NMR Spectroscopy (10 hrs)

Nuclear Magnetic Resonance Spectroscopy: Theory, instrumentation and application to the structural determination of organic compounds (simple organic molecule).

UNIT –4: Mass Spectrometry**(10 hrs)**

Mass Spectrometry: Principle, details of components of mass spectrometer, various analyzers, principles of fragmentation and applications to the organic molecules for structure identification.

UNIT –5: Radio - Analytical Techniques and Computers in Chemistry (10 hrs)

Radio-Analytical technique: Tracer technique, radiocarbon dating, activation analysis, radiometric analysis and titrations and isotopic dilution analysis.

Essentials of computers in chemistry, Fundamentals and applications in chemistry (Data interpretation – softwares for basic chemical calculations and chemical structures).

TEXT BOOKS

1. R Gopalan, K. Rengarajan, P.S. Subramanian, Elements of Analytical Chemistry, Sultan Chand & Sons, third edition, 2003.
2. David Harvey, Modern Analytical Chemistry, McGraw-Hill, first edition, 2000.
3. R. C. Mendham, J. D. Denney, and Barnes M. Thomas, Vogel's Text Book of Quantitative Chemical Analysis, Pearson Education Pvt. Ltd., sixth edition, 2004.
4. K. V. Raman, Computers in Chemistry, Tata McGraw-Hill Education, first edition, 1993.

REFERENCE BOOKS

1. Douglas A. Skoog, Donald M. West and F. James Holler, Fundamentals of Analytical Chemistry, Harcourt Asia Pvt. Ltd., ninth edition, 2001.
2. Douglas A. Skoog, Donald M. West and F. James Holler, Analytical Chemistry- An Introduction, Saunders College Publishers, seventh edition, 2000.
3. Dean, John A., Merritt, Lynne L., Settle, Frank A., Willard, Hobart H; Instrumental Methods of Analysis, Wadsworth Publishing Co Inc., seventh edition, 1988.

4. D. A. Skoog, Principles of Instrumental Analysis, Saunders College Publishing, Philadelphia, London, fifth edition, 1998.

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	Unit – 4	1	
	Unit - 5	1	

CORE PRACTICAL – 3
GRAVIMETRIC ANALYSIS

SUBJECT CODE: 19UCHE311P	PRACTICAL	MARKS 100
SEMESTER: V	CREDITS: 3	TOTAL HOURS: 45

Objectives

To develop experimental skill in the gravimetric analysis of inorganic elements.

Gravimetric Estimation:

1. Estimation of Lead as Lead chromate
2. Estimation of Barium as Barium chromate
3. Estimation of Nickel as Nickel-DMG complex
4. Estimation of Calcium as Calcium oxalate
5. Estimation of Barium as Barium sulfate
6. Estimation of sulfate as Barium sulfate
7. Estimation of lead in solder by gravimetry*
8. Estimation of Nickel from stainless steel>(* For Internal assessment only)

REFERENCE BOOKS

1. V.Venkateswaran, R.Veerarwamy and A.R. Kulandaivelu, Basic Principles of Practical Chemistry, Sultan Chand & Sons, New Delhi, second edition, 1997.
2. B. S. Furniss, Vogel's Textbook of Practical Organic Chemistry, ELBS - Longman, London, seventh edition, 1984.

CORE PRACTICAL – 4
ANALYSIS AND PREPARATION OF ORGANIC COMPOUNDS

SUBJECT CODE: 19UCHE312P	THEORY	MARKS 100
SEMESTER: V	CREDITS: 3	TOTAL HOURS: 45

Objectives

To exercise the laboratory skill on analysis of organic compounds and preparation of organic compounds.

A. Organic Analysis

- a) Qualitative analysis of simple organic compounds:
- b) Confirmation by preparation of solids derivatives / characteristics colour reaction.

Note :1. Mono - functional compounds are given for analysis. In the case of bi-functional compounds, students are required to report any one of the functional groups.

2. Each student is expected to do the analysis of at least 15 different organic substances.

B. Organic Preparations and Chromatography Techniques

- 1. Preparation of Organic compounds involving the following chemical conversions: i.Oxidation ii.Reduction iii.Esterification iv.Acetylation v.Hydrolysis vi.Nitration vii.Bromination viii.Diazotization ix.Osazone formation.
- 2. Separation of chlorophyll from the plant extract by column chromatography*.
- 3. Identification of Food colorants by TLC.(* For Internal assessment only)

REFERENCE BOOKS

1. V.Venkateswaran, R.Veerawamy and A.R. Kulandaivelu, Basic Principles of Practical Chemistry, Sultan Chand & Sons, New Delhi, second edition, 1997.
2. B. S. Furniss, Vogel's Textbook of Practical Organic Chemistry, ELBS - Longman, London, seventh edition, 1984.

CORE PRACTICAL – 5
PHYSICAL CHEMISTRY PRACTICAL

SUBJECT CODE: 19UCHE313P	PRACTICAL	MARKS 100
SEMESTER: V	CREDITS: 3	TOTAL HOURS: 45

Objectives:

To perform the laboratory experiments in order to understand the concepts of physical changes in chemistry.

List of Experiments

1. Critical Solution Temperature
2. Effect of Electrolyte on Critical Solution Temperature
3. Transition Temperature
4. Kinetics of Ester Hydrolysis
5. Kinetics of Persulphate - Iodide Reaction.
6. Rast Method
7. Phase Diagram (Simple eutectic System)
8. Conductometric : Acid - Base Titration
9. Potentiometric: Acid - Base Titration

REFERENCE BOOKS

1. V.Venkateswaran, R.Veerawamy and A.R.Kulandaivelu, Basic Principles of Practical Chemistry, Sultan Chand & Sons, New Delhi, second edition, 1997.
2. Daniel, Experimental Physical Chemistry, McGraw-Hill, New York, seventh edition, 1970.
3. A. Findlay, Practical Physical Chemistry, Logman, London, seventh edition, 1959.
4. V. K. Ahluwalia, Sunith Dhingra, and Adarsh. Gulati, College Practical Chemistry, Orient Longman Pvt. Ltd., 2008.
5. K. K. Sharma and D. S. Sharma, Introduction to Practical Chemistry, Vikas Publishing House, New Delhi, 1982.

ELECTIVE-1
NANOCHEMISTRY AND TECHNOLOGY

SUBJECT CODE: 19UCHE303	THEORY	MARKS 100
SEMESTER: VI	CREDITS: 5	TOTAL HOURS: 75

Objectives:

To create general awareness on nanomaterials and applications.

UNIT - 1: Background to Nanotechnology **(10 hrs)**

Origin of nanotechnology - scientific revolution - atomic structures - molecular and atomic size - Bohr radius - Heisenberg's principle. Definitions - Nano, nanoscience, and nanotechnology, nanograins, nanoclusters, nanoinclusions (Lycurgus cup etc). Nano in nature - difference between bulk and nanomaterials - challenges in nanotechnology - carbon age - new form of carbon (from Graphene sheet to CNT).

UNIT - 2: Preparation and Properties of Nanomaterials – I **(15 hrs)**

Influence of nucleation rate on the size of the crystals - macroscopic to microscopic crystals and nanocrystals - large surface to volume ratio, top-down and bottom-up approaches, self-assembly process - grain boundary volume in nanocrystals - defects in nanocrystals - surface effects on the properties. Size dependent properties - magnetic, electronic, transport and optical.

UNIT - 3: Preparation and Properties of Nanomaterials – II **(20 hrs)**

Synthesis of bulk nanostructured materials - Sol gel processing - mechanical alloying and milling - inert gas condensation technique - bulk and nano composite materials – grinding high energy ball milling - types of balls - WC and ZrO₂- materials – ball ratio – limitations melt quenching and annealing. Preparation and properties of Carbon Nanotubes (CNT) - Metals (Au, Ag) - Metal oxides (TiO₂, ZnO) - Semiconductors (Si, CdS) - nanocomposites - Dilute magnetic semiconductor. Organic nanomaterials – dendrimers, etc.

UNIT - 4: Applications of Nanotechnology

(20 hrs)

Applications - electronics, sensors, catalysis, nanocomposites (NCMs) and environmental. Current medical practice -treatment methodology- principles of nanomedicine – nanomedical perspective and the medical applications – Nanomedicine: diagnosis, nanopharmaceuticals, biocompatible nanomedical materials. Industrial applications of nanomaterials: nanocoatings and nanotextiles as antibacterial and anti – odour agents in deodorant/antiperspirant, shaving/depilatory products, foot powder, oral care.

UNIT –5: Nanotoxicity

(10 hrs)

Ethical, safety and regulatory issues of nanomedicine. Nano toxicology: toxicity of carbon nanomaterials, handling of nanomaterials, health implication of nanomaterials, environmental toxicity. Green Nanochemistry.

REFERENCE BOOKS

1. M. Wilson, K. Kannangara, G Smith, M. Simmons, B. Raguse, Nanotechnology: Basic Science and Emerging Technologies, Overseas Press India Pvt. Ltd., New Delhi, first edition, 2005.
2. C. N. R. Rao, A. Muller, A. K. Cheetham , The Chemistry of Nanomaterials Synthesis, Properties and Applications, Wiley VCH Verlag GmbH & Co., Weinheim, 2004.
3. Kenneth J. Klabunde , Nanoscale Materials Science, John Wiley & Sons, Inc., 2001.
4. C. S. S. R. Kumar, J. Hormes, C. Leuschner, Nanofabrication Towards Biomedical Applications, Wiley –VCH Verlag GmbH & Co, Weinheim, 2004.
5. W. Rainer, Nano Electronics and information Technology, Wiley, 2003.
6. K. E. Drexler, Nano systems, Wiley, 1992.
7. G. Cao, Nanostructures and Nanomaterials: Synthesis, Properties and Applications, Imperial College Press, 2004.
8. M. Madou, Fundamentals of Micro-fabrication, CRC press, 1997.
9. G. Timp, Nanotechnology, AIP press, Springer Verlag, New York, 1999.

10. M. J. Jackson, Micro fabrication and Nanomanufacturing, CRC press. 2005.
11. G. Cao, Nanostructures and Nanomaterials: Synthesis, Properties and Applications, Imperial College Press, 2004
12. W. T. S Huck, Nanoscale assembly: Chemical Techniques (Nanostructure Science and Technology, Springer, 2005.
13. Robert A. Freitas Jr., Nanomedicine, Volume I: Basic Capabilities, Landes Bioscience, Georgetown, TX, 1999.
14. Robert A. Freitas Jr., Nanomedicine, Volume IIA: Biocompatibility, Landes Bioscience, Georgetown, TX, 2003.
15. Kewal K. Jain, The Hand Book of Nanomedicine, Humana Press, Springer 2008.
16. Challa S. S. R. Kumar, Nanomaterials for Medical Diagnosis and Therapy, Wiley - VCH, 2007.
17. Parag Diwan and Ashish Bharadwaj, Nano Medicines, Pentagon Press, 2006.
18. Nancy A. Monteiro - Riviere and C. Lang Tran, Nanotoxicology, Characterization, Dosing and Health Effects, Informa Healthcare, 2007.
19. Kumar, Challa S. S. R., Nanomaterials - Toxicity, Health and Environmental Issues, Wiley - VCH, Weinheim, 2006.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
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Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ELECTIVE-2
INDUSTRIAL CHEMISTRY

SUBJECT CODE: 19UCHE318	THEORY	MARKS 100
SEMESTER: II	CREDITS: 5	TOTAL HOURS: 75

Objectives

To have the thorough knowledge on various chemical industries and their manufacturing processes.

UNIT - 1: Industrial Requirements **(15 hrs)**

Requirements of an industry - location - water - industrial water treatment - safety measures - pilot plants. Fuels - types of fuels with examples - coal - carbonization of coal - coal tar distillation - liquid fuels - gaseous fuels - selection of fuels - nuclear fuels. Energy - sources of energy - renewable and non-renewable energies - non conventional energies. Industrial catalysts - Types of catalysts - Functions and applications of Raney Nickel, Pd, CuCrO₄, TiO₂, Al, V and Pt based catalysts and zeolites.

UNIT - 2: Petrochemical Industries **(15 hrs)**

Crude oil - constitution and distillation - composition of different distillates - pour points, depressants, drag reducers, viscosity reducers, ignition point, flash point octane number - cracking - catalysts used in petroleum industries - structure, selectivity and applications. Manufacture of synthetic petrol - Dergius and Fischer Tropsh processes - Manufacture of petrochemicals and petrochemical polymers - Manufacture of higher olefins, Acetaldehyde, Acetic acid, Ethylene glycol, Glycerine, Acetone, Phenol, Carbon disulphide, Vinylacetate, Cumene, Chlorophrene, Butane diols, Xylenes.

UNIT - 3: Fertilizers and Speciality Chemicals **(15 hrs)**

Manufacture - Properties and industrial uses of solvents - DMF, DMSO, THF and Dioxane. Fertilizers - Raw materials, manufacture (flow chart chemical process with equations) of ammonium nitrate, ammonium sulphate, urea, calcium cyanamide, calcium ammonium nitrate, sodium nitrate,

ammonium chloride, ammonium phosphate, super phosphate of lime, NPK fertilizers, Manufacture in pure form of the following – Sodium carbonate, Oxalic acid, Potassium dichromate, Perchloric acid.

UNIT - 4: Oils, Soaps and Detergents (15 hrs)

Manufacture of Cl_2 , NaOH and Chlorates of Na and K - manufacture of perchlorate. Oils - difference between oils and fats - manufacture of cotton seed oil and soybean oil - refining of oil - manufacture of soaps - toilet and transparent soaps - Detergents - synthetic detergents - surface active agents and their classification - manufacture of anionic, cationic and non ionic detergents and shampoo.

Sugar industry - manufacture of sugar from cane sugar and beet root. Manufacture of leather - hides - Vegetable and chrome tanning finishing. Manufacture of dinitrophenols, malathion, parathion, schradan and dementon.

UNIT - 5: Metallurgy (15 hrs)

General methods of metallurgy - ores - types - methods of concentration of ores - hydro metallurgy, pyrometallurgy - various reduction process, refining of metals - extraction of Cr, Mn, V, Co, Pt, U and Th. Environmental problems of chemicals industries - methods of control - sewage treatment and waste management.

REFERENCEBOOKS

1. B.K. Sharma, Industrial Chemistry, Goel Publishing House, Meerut, 2003
2. C. E. Dryden, Outlines of Chemical Technology, Gopala Rao, Eastwest Press, New Delhi, third edition, 1997.
3. G.T. Austin, R.N. Shreve, Chemical Process Industries, Tata McGraw Hill publishing company, Mumbai, fifth edition, 1984.
4. H. Steines, Introduction to Petrochemicals, Pergamon Press, 1961.
5. Alan Cottrell, An Introduction to Metallurgy, Orient Longman, 2000.

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	Unit – 3	1	
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Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ELECTIVE-3
PHARMACEUTICAL CHEMISTRY

SUBJECT CODE: 19UCHE317	THEORY	MARKS 100
SEMESTER: VI	CREDITS: 5	TOTAL HOURS: 75

Objectives

- *To understand the pharmacology, drugs and their mechanism of action, various diseases and their cure.*
- *To introduce the knowledge on pharmaceutical industries and their functioning.*

UNIT - 1: Introduction **(15 hrs)**

Common diseases – Infective diseases – insect - born, air - born and water - born – hereditary diseases. Terminology of drugs, pharmacology, pharmacognesy, pharmacodynamics, pharmacokinetics, antimetabolites. Indian medicinal plants – Tulsi, Neem, Keezhanelli and their importance.

UNIT - 2: Pharmacology and Metabolism **(15 hrs)**

Mechanism of drug action – Action at cellular and extra cellular sites. Absorption of drugs – routes of administration , factors affecting absorption – Assay of drugs – chemical, biological, immunological assays, LD₅₀ and ED₅₀, therapeutic index, drug dosage. Metabolism of drugs through oxidation, reduction, hydrolysis and conjugate processes; factors affecting metabolism.

UNIT - 3: Drugs **(15 hrs)**

Definition and two examples each: Anaesthetics – General and local; Analgesics – Narcotic and synthetic; Antipyretics and anti inflammatory agents. Antibiotics: penicillin, streptomycin, chloramphenicol and tetracyclins – Antivirals, AIDS: symptoms, prevention, treatment – Cancer and neoplastic agents.

UNIT - 4: Common Body Ailments **(15 hrs)**

Diabetes – Causes, hyper and hypoglycemic drugs – Blood pressure –

Systolic and Diastolic, Hypertensive drugs – cardiovascular drugs – antiarrhythmic, antianginals, vasodialators – Psychedelic drugs, hypnotics, sedatives – Lipid profile – HDL, LDL cholesterol, lipid lowering drugs. Composition of Blood: Blood grouping and Matching – Role of blood as Oxygen carrier, Coagulation of Blood - Coagulants, Anticoagulants. Anaemia – Causes and control.

UNIT - 5: Pharma Industries: Functioning and Documentations (15 hrs)

Introduction of pharma industry - research and development unit - process development and batch manufacturing - quality control and quality assurance – specifications Standard Operating Procedures (SOP) – Standard Testing Procedures (STP) – Material safety data sheets– journals – pharmacopeia – patents (product and process) – reports and records. Regulatory affairs for pharma industries like GLP, GMP, ISO, and FDA.

REFERENCE BOOKS

1. Jayashree Ghosh, Pharmaceutical Chemistry, S. Chand and Company Ltd.,
2. S. Lakshmi, Pharmaceutical Chemistry S. Chand & Sons New Delhi, third edition, 1995.
3. Ashutosh Kar, Medicinal Chemistry, New Age International Ltd., New Delhi, 1993.
4. Divid A. William, Thomas Lemke, O. Foyes, Principles of Medicinal Chemistry, Bpublishers, fifth Edition, 2005.
5. Bertram G. Katzung, Basic & Clinical Pharmacology, McGraw-Hill Medical, twelthedition, 2012.
6. E. Prichard, Quality in the Analytical Chemistry Laboratory, John Wiley and sons New York, 1997.
7. W Funk, V Dammann, G. Donnevert, Quality Assurance in Analytical, VCH Weinheim, New York, 1995.

New Delhi, f

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

**ELECTIVE-4:
APPLIED ELECTROCHEMISTRY**

SUBJECT CODE:	THEORY	MARKS 100
SEMESTER: VI	CREDITS: 5	TOTAL HOURS: 75

Objectives

To know the applications of electrochemistry in industrial processes viz., metallurgy, coating, cells and corrosion.

UNIT - 1: Industrial Electrochemistry (15 hrs)

Electrochemical process in industry - components of electrochemical reactions - Types of electrolytes - Cathodes and anodes in electrochemical reactor - separators. Inorganic electrochemical: Caustic soda and chlorine production, mercury cells, diaphragm cells, membrane cells - Advantage chlorates, perchlorates, hydrogen peroxide. Organic electrochemical: Special feature of electro-organic synthesis - electro chemical oxidation. Kolbe synthesis - electro reduction of carbonyl compounds - adiponitrile synthesis.

UNIT - 2: Electrometallurgy (15 hrs)

Electro deposition of metals - principles - nucleation and growth of crystals - nature of electro deposits. Hydrometallurgy: Recovery of meals. Recovery of silver from photographic emulsion - electro refining - production of high copper; process description. Pyro-metallurgy: Necessity for using molten electrolytes - reactors for molten salt electrolytes - Production of aluminium - electrodes and electrode reactions in cryolite melt. Electrochemical purification of aluminium - other metals through molten salt electrolysis - Magnesium and sodium brief outline.

UNIT 3: Electroplating (15 hrs)

Fundamental principles - nature of deposits for electroplating - Hull cell experiments - operating conditions and nature of deposits - throwing power - preparation of samples for electroplating - chemical and electrochemical cleaning

- electroplating of copper, nickel and cadmium. Electroless plating: Importance, plating of non-metals - both composition - electroless plating of copper and nickel.

UNIT - 4: Electrochemical Power Sources

(15 hrs)

Basic principles, chemical and electrical energies - inter conversion - changing and discharging - requirements for a good power source - Types of power sources. Primary Batteries: Description of primary cells - alkaline, manganese cells - silver oxide - zinc cells - lithium primary cells - applications. Secondary Batteries: Importance applications - charge discharge efficiency - cycle life - energy density - lead acid batteries for electric vehicles. Fuel Cells: Basic principles - Hydrogen, oxygen fuel cells - gas diffusion electrodes for fuel cells - alkaline fuel cells.

UNIT - 5: Corrosion and Prevention

(15 hrs)

Principles - stability of metals - EMF series - active and noble metals. pH effect of stability - Pourbaix diagram - Kinetics of corrosion - mixed potential process - cathodic reaction - anodic reaction - corrosion current - active dissolution - passivation - break down of passivity - Evans diagram. Method of corrosion protection. Principles and inhibition of anodic, cathodic processes - Inhibitive additives for corrosion protection - protective coatings- types of coatings - protection of structures and pipelines - cathodic protection - examples - sacrificial anodes - protection of ships in sea water.

REFERENCE BOOKS

1. D. Pletcher and F.C. Waish, Industrial Chemistry, second edition, 1990.
2. C.H. Hamann, A. Hamnett and W. Vielstich, Electrochemistry, Weinheim-WileyVCH, 1998.
3. D.B. Hibbert, Introduction to Electrochemistry, McMillan, London, 1993.

Question paper pattern:

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
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Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ELECTIVE – 5
POLYMER CHEMISTRY

SUBJECT CODE: 19UCHE319	THEORY	MARKS 100
SEMESTER: VI	CREDITS: 5	TOTAL HOURS: 75

Objectives:

To develop the knowledge on chemistry of polymers and their applications.

UNIT - 1: Introduction to Polymers (15 hrs)

Importance of polymers: Basic concept - Monomers and polymers - definition. Classification of polymers on the basis of microstructures, macrostructures and applications (thermosetting and thermoplastics). Distinction among plastics, elastomers and fibers. Homo and hetero polymers and copolymers.

Chemistry of polymerization: Chain polymerisation, free radical, ionic, coordination and step polymerisation. Polyaddition and polycondensation - miscellaneous ring - opening and group transfer polymerisations.

UNIT - 2: Physical Properties and Reactions of Polymers (15 hrs)

Properties: Glass transition temperature (T_g) - Definition - Factors affecting T_g - relationships between T_g and molecular weight and melting point. Importance of T_g . Molecular weight of polymers: Number average, weight average, sedimentation and viscosity average molecular weights. Molecular weights and degree of polymerisation. Reactions: hydrolysis - hydrogenation - addition - substitutions - cross-linking, vulcanisation and cyclisations reactions. Polymer degradation: Basic idea of thermal, photo and oxidative degradations of polymers.

UNIT - 3: Polymerization Techniques and Processing (15 hrs)

Polymerisation techniques: Bulk, solution, suspension, emulsion, melt condensation and interfacial polycondensation polymerisations. Polymer processing: Calendering - die casting, rotational casting - compression. Injection moulding.

UNIT - 4: Chemistry of Commercial Polymers**(15 hrs)**

General methods of preparation, properties and uses of the following: Teflon, polymethylmethacrylate. Polythylene, polystyrene, PAN, polyesters, polycarbonates, polyamides (Kevlar), polyurethanes, PVC, epoxy resins, rubber styrene and neoprene rubbers, Phenol-formaldehydes and urea - formaldehyde resins.

UNIT - 5: Advances in Polymers**(15 hrs)**

Biopolymers – biomaterials; polymers in medical field. High temperature and fire - resistant polymers – Silicones - carbon Fibers (Basic idea only).

TEXT BOOK

1. F.W. Billmeyer, Text book of Polymer Science, Jr. John Wiley & Sons, 1984.

REFERENCE BOOKS

1. V.R. Gowariker, N.V. Viswanathan and Jayader Sreedhar, Polymer Science, WileyEastern Ltd., New Delhi, 1978.
2. B.K. Sharma, Polymer Chemistry, Goel Publishing House, Meerut, 1989.
3. M.G. Arora, and M.S. Yadav, Polymer Chemistry, Anmol Publications Private Ltd., NewDelhi, second revised edition,1989.

Question paper pattern

Section	Question Component	Numbers	Marks	Total
Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
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TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
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	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
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	Unit – 3	1	
	Unit – 4	2	
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Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	



ALLIED CHEMISTRY

**(For B.Sc., Physics,
Mathematics,
Plant Biology & Plant
Biotechnology and
Advanced Zoology &
Biotechnology)**

SYLLABUS FOR ALLIED CHEMISTRY

ALLIED CHEMISTRY - I (For Mathematics & Physics Students)

SUBJECT CODE: 19UCHE331	THEORY	MARKS 100
SEMESTER: I	CREDITS: 3	TOTAL HOURS: 45

UNIT – 1: Chemical Bonding and Nuclear Chemistry

Chemical Bonding: Molecular Orbital Theory-bonding, antibonding and nonbonding orbitals. Molecular orbital diagrams for Hydrogen, Helium, Nitrogen; discussion of bond order and magnetic properties.

Nuclear Chemistry: Fundamental particles - Isotopes, Isobars, Isotones and Isomers- Differences between chemical reactions and nuclear reactions - group displacement law. Nuclear binding energy - mass defect - calculations. Nuclear fission and nuclear fusion - differences – Stellar energy. Applications of radioisotopes - carbon dating, rock dating and medicinal applications.

UNIT – 2: Industrial Chemistry

Fuels: Fuel gases: Natural gas, water gas, semiwater gas, carbureted water gas, producer gas, CNG, LPG and oil gas (manufacturing details not required). Silicones: Synthesis, properties and uses of silicones.

Fertilizers: Urea, ammonium sulphate, potassium nitrate, NPK fertilizer, super phosphate, triple superphosphate.

UNIT – 3: Fundamental Concepts in Organic Chemistry

Hybridization: Orbital overlap, hybridization and geometry of CH₄, C₂H₄, C₂H₂ and C₆H₆. Electronic effects: Inductive effect and consequences on K_a and K_b of organic acids and bases, electromeric, mesomeric, hyperconjugation and steric- examples and explanation.

Reaction mechanisms: Types of reactions - aromaticity- aromatic electrophilic substitution; nitration, halogenation, Friedel - Craft's alkylation and acylation.

Heterocyclic compounds: Preparation, properties of pyrrole and pyridine.

UNIT - 4: Thermodynamics and Phase Equilibria

Thermodynamics: Types of systems, reversible and irreversible processes, isothermal and adiabatic processes and spontaneous processes. Statements of first law and second law of thermodynamics. Carnot's cycle and efficiency of heat engine. Entropy and its significance. Free energy change and its importance (no derivation). Conditions for spontaneity in terms of entropy and Gibbs free energy. Relationship between Gibbs free energy and entropy.

Phase Equilibria: Phase rule - definition of terms in it. Applications of phase rule to water system. Two component system - Reduced phase rule and its application to a simple eutectic system (Pb-Ag).

UNIT – 5: Analytical Chemistry

Introduction to qualitative and quantitative analysis. Principles of volumetric analysis. Separation and purification techniques – extraction, distillation and crystallization.

Chromatography: principle and application of column, paper and thin layer chromatography.

REFERENCE BOOKS

1. V. Veeraiyan, Text book of Ancillary Chemistry; Highmount publishing house, Chennai, first edition, 2009.
2. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006.
3. Arun Bahl, B. S. Bahl, Advanced Organic Chemistry; S. Chand and Company, New Delhi, twenty third edition, 2012.
4. P. L. Soni, H. M. Chawla, Text Book of Organic Chemistry; Sultan Chand & sons, New Delhi, twenty ninth edition, 2007.
5. P. L. Soni, Mohan Katyal, Text book of Inorganic chemistry; Sultan Chand and Company, New Delhi, twentieth edition, 2007.
6. B. R. Puri, L. R. Sharma, M. S. Pathania, Text book Physical Chemistry; Vishal Publishing Co., New Delhi, forty fourth edition, 2010.
7. B. K, Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014.

Question paper pattern

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Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
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TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
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	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ALLIED CHEMISTRY - II
(For Mathematics & Physics Students)

SUBJECT CODE: 19UCHE335	THEORY	MARKS 100
SEMESTER: II	CREDITS: 3	TOTAL HOURS: 45

UNIT – 1: Co-ordination Chemistry and Water Technology

Co-ordination Chemistry: Definition of terms - IUPAC Nomenclature - Werner's theory - EAN rule - Pauling's theory – Postulates - Applications to $[\text{Ni}(\text{CO})_4]$, $[\text{Ni}(\text{CN})_4]^{2-}$, $[\text{Co}(\text{CN})_6]^{3-}$ Chelation - Biological role of Haemoglobin and Chlorophyll (elementary idea) - Applications in qualitative and quantitative analysis.

Water Technology: Hardness of water, determination of hardness of water using EDTA method, zeolite method - Purification techniques - BOD, COD.

UNIT – 2: Carbohydrates and Amino acids

Carbohydrates: Classification, preparation and properties of glucose, fructose and sucrose. Discussion of open chain ring structures of glucose and fructose. Glucose –fructose interconversion. Properties of starch and cellulose.

Amino acids: Classification - preparation and properties of alanine, preparation of dipeptides using Bergmann method. RNA and DNA (elementary idea only).

UNIT – 3: Electrochemistry

Galvanic cells - Standard hydrogen electrode - calomel electrode - standard electrode potentials - electrochemical series. Strong and weak electrolytes - ionic product of water - pH, pKa, pKb. Conductometric titrations - pH determination by colorimetric method - buffer solutions and its biological applications - electroplating - Nickel and chrome plating - Types of cells - fuel cells - corrosion and its prevention.

UNIT – 4: Kinetics and Catalysis

Order and molecularity. Integrated rate expression for I and II ($2A \rightarrow$ Products) order reactions. Pseudo first order reaction, methods of determining order of a reaction - Half life period – Catalysis - homogeneous and heterogeneous, catalyst used in Contact and Haber's processes. Concept of energy of activation and Arrhenius equation.

UNIT – 5: Photochemistry

Grothus-Drapper's law and Stark-Einstein's law of photochemical equivalence, Quantum yield - Hydrogen-chloride reaction. Phosphorescence, fluorescence, chemiluminescence and photosensitization and photosynthesis (definition with examples).

REFERENCE BOOKS

1. V. Veeraiyan, Text book of Ancillary Chemistry; Highmount publishing house, Chennai, first edition, 2009.
2. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006.
3. Arun Bahl, B. S. Bahl, Advanced Organic Chemistry; S. Chand and Company, New Delhi, twenty third edition, 2012.
4. P. L. Soni, H. M. Chawla, Text Book of Organic Chemistry; Sultan Chand & sons, New Delhi, twenty ninth edition, 2007.
5. P. L. Soni, Mohan Katyal, Text book of Inorganic chemistry; Sultan Chand and Company, New Delhi, twentieth edition, 2007.
6. B. R. Puri, L. R. Sharma, M. S. Pathania, Text book Physical Chemistry; Vishal Publishing Co., New Delhi, forty fourth edition, 2010.
7. B. K, Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014.

Question paper pattern

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Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ALLIED CHEMISTRY - I
(For Botany and Zoology Students)

SUBJECT CODE: 19UCHE336	THEORY	MARKS 100
SEMESTER: I	CREDITS: 3	TOTAL HOURS: 45

UNIT – 1: Chemical Bonding and Nuclear Chemistry

Chemical Bonding: Molecular Orbital Theory-bonding, antibonding and nonbonding orbitals. M. O diagrams for Hydrogen, Helium, Nitrogen; discussion of bond order and magnetic properties.

Nuclear Chemistry: Fundamental particles - Isotopes, Isobars, Isotones and Isomers- Differences between chemical reactions and nuclear reactions - group displacement law. Nuclear binding energy - mass defect - calculations. Nuclear fission and nuclear fusion - differences – Stellar energy. Applications of radioisotopes - carbon dating, rock dating and medicinal applications.

UNIT – 2: Industrial Chemistry

Fuels: Fuel gases: Natural gas, water gas, semiwater gas, carbureted water gas, producer gas, CNG, LPG and oil gas (manufacturing details not required).

Silicones: Synthesis, properties and uses of silicones.

Fertilizers: Urea, ammonium sulphate, potassium nitrate NPK fertilizer, superphosphate, triple superphosphate.

UNIT – 3: Fundamental Concepts in Organic Chemistry

Hybridization: Orbital overlap hybridization and geometry of CH₄, C₂H₄, C₂H₂ and C₆H₆. Polar effects: Inductive effect and consequences on K_a and K_b of organic acids and bases, electromeric, mesomeric, hyperconjugation and steric- examples and explanation.

Reaction mechanisms: Types of reactions - aromaticity- aromatic electrophilic substitution; nitration, halogenation, Friedel - Craft's alkylation and acylation.

Heterocyclic compounds: Preparation, properties of pyrrole and pyridine.

UNIT – 4: Useful Organic compounds

Definition, structure and uses: Antibiotics viz., Penicillin, Chloramphenicol and Streptomycin; Anaesthetics viz., Chloroform and ether; Antipyretics viz., aspirin, paracetamol and ibuprofen; Artificial sugar viz., saccharin, Aspartame and cyclamate; Organic Halogen compounds viz., Freon, Teflon.

UNIT – 5: Analytical Chemistry

Introduction qualitative and quantitative analysis. Principles of volumetric analysis. Separation and purification techniques: extraction, distillation and crystallization. Chromatography: principle and application of column, paper and thin layer chromatography.

REFERENCE BOOKS

1. V. Veeraiyan, Text book of Ancillary Chemistry; Highmount publishing house, Chennai, first edition, 2009.
2. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006.
3. Arun Bahl, B. S. Bahl, Advanced Organic Chemistry; S. Chand and Company, New Delhi, twenty third edition, 2012.
4. P. L. Soni, H. M. Chawla, Text Book of Organic Chemistry; Sultan Chand & sons, New Delhi, twenty ninth edition, 2007.
5. P. L. Soni, Mohan Katyal, Text book of Inorganic chemistry; Sultan Chand and Company, New Delhi, twentieth edition, 2007.
6. B. R. Puri, L. R. Sharma, M. S. Pathania, Text book Physical Chemistry; Vishal Publishing Co., New Delhi, forty fourth edition, 2010.
7. B. K, Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014.
8. Jayashree gosh, Fundamental Concepts of Applied Chemistry; Sulan & Chand, Edition – 2006.

Question paper pattern

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Section B	Short Answer Answer any 5 out of 7 questions	13–19	6	30
Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ALLIED CHEMISTRY - II
(For Botany and Zoology Students)

SUBJECT CODE: 19UCHE338	THEORY	MARKS 100
SEMESTER: II	CREDITS: 3	TOTAL HOURS: 45

UNIT – 1: Co-ordination Chemistry and Water Technology

Co-ordination Chemistry: Definition of terms - IUPAC Nomenclature - Werner's theory - EAN rule - Pauling's theory – Postulates - Applications to $[\text{Ni}(\text{CO})_4]$, $[\text{Ni}(\text{CN})_4]^{2-}$, $[\text{Co}(\text{CN})_6]^{3-}$ Chelation - Biological role of Haemoglobin and Chlorophyll (elementary idea) - Applications in qualitative and quantitative analysis.

Water Technology: Hardness of water, determination of hardness of water using EDTA method, zeolite method - Purification techniques – BOD and COD.

UNIT – 2: Carbohydrates

Classification, preparation and properties of glucose and fructose. Discussion of open chain ring structures of glucose and fructose. Glucose-fructose interconversion. Preparation and properties of sucrose, starch and cellulose.

UNIT – 3: Amino Acids and Essential elements of biosystem

Classification - preparation and properties of alanine, preparation of dipeptides using Bergmann method - Proteins-classification – structure - Colour reactions - Biological functions –nucleosides - nucleotides - RNA and DNA – structure. Essentials of trace metals in biological system- Na, Cu, K, Zn, Fe, Mg.

UNIT – 4: Electrochemistry

Galvanic cells - Standard hydrogen electrode - calomel electrode - standard electrode potentials - electrochemical series. Strong and weak electrolytes - ionic product of water - pH, pKa, pKb. Conductometric titrations - pH determination by colorimetric method - buffer solutions and its biological applications - electroplating - Nickel and chrome plating - Types of cells - fuel

cells - corrosion and its prevention.

UNIT – 5: Photochemistry

Grothus-Drapper's law and Stark-Einstein's law of photochemical equivalence, Quantum yield - Hydrogen-chloride reaction. Phosphorescence, fluorescence, chemiluminescence and photosensitization and photosynthesis (definition with examples).

REFERENCE BOOKS

1. V. Veeraiyan, Text book of Ancillary Chemistry; Highmount publishing house, Chennai, first edition, 2009.
2. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006.
3. Arun Bahl, B. S. Bahl, Advanced Organic Chemistry; S. Chand and Company, New Delhi, twenty third edition, 2012.
4. P. L. Soni, H. M. Chawla, Text Book of Organic Chemistry; Sultan Chand & sons, New Delhi, twenty ninth edition, 2007.
5. P. L. Soni, Mohan Katyal, Text book of Inorganic chemistry; Sultan Chand and Company, New Delhi, twentieth edition, 2007.
6. B. R. Puri, L. R. Sharma, M. S. Pathania, Text book Physical Chemistry; Vishal Publishing Co., New Delhi, forty fourth edition, 2010.
7. B. K, Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014.

Question paper pattern

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Section C	Essay Answer any 4 out of 6 questions	20– 25	10	40
TOTAL MARKS				100

Distribution of Questions

Sections	Units	No. of Questions	
		Theory	Problems
Section A	Unit – 1	3	
	Unit – 2	3	
	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
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	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

ALLIED CHEMISTRY PRACTICAL
(Common for Mathematics, Physics, Botany and Zoology students)

SUBJECT CODE: 19UCHE332P / 19UCHE337P	THEORY	MARKS 100
SEMESTER: ODD & EVEN	CREDITS: 4	TOTAL HOURS: 60

1. VOLUMETRIC ANALYSIS

1. Estimation of Sodium hydroxide using standard Sodium Carbonate.
2. Estimation of Hydrochloric acid using standard Oxalic acid.
3. Estimation of Ferrous sulphate using standard Mohr's salt
4. Estimation oxalic acid using standard Ferrous Sulphate.
5. Estimation of Potassium permanganate using standard Sodium hydroxide.
6. Estimation of Magnesium using EDTA.
7. Estimation of Ferrous ion using diphenylamine as internal indicator.

2. SYSTEMATIC ANALYSIS OF ORGANIC COMPOUNDS

The analysis must be carried out as follows:

- a) Functional group tests [phenol, acids (mono & di) aromatic primary amine, amides(mono & di), aldehyde and glucose].
- b) Detection of elements (N, S, Halogens)
- c) Distinguish between aliphatic and aromatic
- d) Saturated and unsaturated compounds.

REFERENCE BOOK

V. Venkateswaran, R. Veerasamy, A. R. Kulandaivelu, Basic Principles of Practical Chemistry; Sultan Chand & sons, second edition, 1997.

Inter-disciplinary Elective

(For Students other than
Chemistry major)(w.e.f. 2019-
2020)

SUBJECT CODE: 19UIDE305	THEORY	MARKS 100
SEMESTER: V	CREDITS: 5	TOTAL HOURS: 60

CHEMISTRY IN EVERYDAY LIFE

Unit – I Dairy chemistry (10 hrs)

- 1.1 Milk Definition, general composition - physico - chemical changes taking place in milk due to boiling, pasteurization, sterilization and homogenisation - explanation.
- 1.2 Components of milk - lipids, proteins, carbohydrates, vitamins, ash and mineral matters - names and functions.
- 1.3 Definition and compositions of cream, butter, ghee, icecream, stabiliser, emulsifier and emulsifier.

Unit – II Food Chemistry (10 hrs)

- 2.1 Sources of food, types, Food adulteration- Common adulterants, Detection of adulterated foods by simple analytical techniques.
- 2.2 Food additives - artificial sweeteners - Saccharin - Cyclamate and Aspartate. Food flavours - esters, aldehydes and heterocyclic compounds. Food colours – Emulsifying agents – preservatives - leavening agents. Baking powder - yeast - taste makers - MSG vinegar
Beverages-soft drinks-soda-fruit juices-alcoholic beverages, examples.

Unit – III Pharmaceutical chemistry (15 hrs)

- 3.1 Definition of the following terms - drug, bacteria, virus and vaccine. Causes, symptoms and drugs for anaemia, jaundice, cholera, malaria, filarial, AIDS, Diabetes, Blood pressure, Cancer.
- 3.2. Indian medicinal plants and uses - tulasi, neem, kizhanelli, mango, semparuthi, adadodai and thoothuvalai.

3.3 Definition and examples for the following – Analgesics, Antipyretics, Anaesthetics, Antibacterials, Antiseptics and Disinfectants (Structures not required)

Unit – IV Industrial Chemistry

(15 hrs)

4.1 Fats and oils - Sources of oils - production of refined vegetable oils - preservation.

4.2 Soaps and Detergents – toilet and transparent soaps-synthetic detergents-surface activeagents and their classification.

4.3 Fertilisers- Types and Importance of fertilizers-. Examples for insecticides, fungicides, pesticides and herbicides - ill effects of use of chemical fertilisers and insecticides.

4.4 Water – sources of water – Hard and soft water-methods of removal of hardness of water.(equations not needed)

Unit- V Polymer Chemistry

(10 hrs)

5.1 Classification and special properties of polymers - natural and synthetic - rubber, cellulose, starch, wool, silk - synthetic rubber, polyalkenes, acrylics, polyamides, polyesters, PVC polyurethane - starting materials and uses only. (structures not required for examination)

5.2 International Resin Identification Coding System and its significance in use - recycling. Bio-degradable polymers. Polymer waste management.

Question paper pattern

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Section A	Definition / Principles Answer any 10 out of 12 questions	1 – 12	3	30
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TOTAL MARKS				100

Distribution of Questions

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	Unit – 3	2	
	Unit – 4	2	
	Unit – 5	2	
Section B	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	2	
	Unit – 5	1	
Section C	Unit – 1	2	
	Unit – 2	1	
	Unit – 3	1	
	Unit – 4	1	
	Unit - 5	1	

CERTIFICATE COURSE –
ENTREPRENEUR SKILLS IN CHEMISTRY

Objectives

To develop entrepreneur skills and to give hands on training to the preparation of the products for the start of small scale industries.

UNIT - 1: Food Chemistry

- Food adulteration - contamination of food items with clay stones, water and toxic chemicals - Common adulterants.
- Food additives, Natural and synthetic anti-oxidants, glazing agents (hazardous effect), food colourants, Preservatives, leavening agents, Baking powder and baking soda, yeast, MSG, vinegar.

UNIT – II: Dairy Chemistry

- Milk - General Composition of milk, constituents of milk- carbohydrates, lipids, proteins, vitamins and minerals. Processing of milk - Pasteurization, Homogenization.
- Special milk - standardized milk, homogenized milk, flavoured milk, toned milk, vitaminized milk, condensed milk-definition, composition and nutritive value.

UNIT- III: Hands on Experience

- Detection of adulterants in food items like coffee, tea, pepper, chilli powder, turmeric powder, butter, ghee, milk, honey etc., by simple techniques.
- Preparation of Jam, squash and Jelly, Gulkand, cottage cheese.
- Preparation of products like candles, soap, detergents, cleaning powder, shampoos, painbalm, tooth paste/powder and disinfectants in small scale.
- Extraction of oils from spices and flowers.
- Testing of water samples using testing kit.

MOOC [Optional]

Students can opt for Massive Open Online Courses during odd semesters. Two credits will be given for each course.

Students can choose a maximum of three courses.
